

A. Metals :- have the following physical properties:

- Good conductors of heat and electricity
- Malleable and ductile, shine
- They are hard and strong Eg. Gold , copper,

NB: mercury is the only metallic liquid at room temperature.

B. Non metals :- are poor conductors

They do not exhibit the properties of metals.

Most of non metals are gases (Cl, O, H, N) or solid (C, S,P)

NB: the only liquid non metal is Bromine

C. Metalloids having both metallic and non metallic properties. Eg, silicon , Germanium, Boron, Antimony, Arsenic

7. 2.2 **Compounds** :- A substance composed of two or more elements chemically combined in a fixed proportion. It can be classified into oxides , bases , acids and salts.

1. Oxides :- formed elements combined with oxygen

Example:- element + Oxygen ➔ Oxide

2. Bases :- that release Hydroxide ions(OH) in water solution. Common bases are sodium Hydroxide (caustic soda), potassium Hydroxide(Caustic potassium), Calcium Hydroxide (Slaked lime)

3. Acids :- that release hydrogen ions(H⁺) in water solution. They have sour tastes

4. Salts:- they can be prepared by into many ways.

I. by neutralization :- by reacting acids and bases

II. by direct reaction of metal and non metals

Exercise

Answer the following questions.

1. How salt is formed

2. List the properties of metals and non metals

3. What is compound?
