questiontext

The _____boiling fractions condense in the lower portion of the tower

The approximate carbon atoms contained in light petrol_____

In Bergius process, the carbon rings in coal undergo _____ to give smaller fragments which are then hydrogenated to open chain and cyclic hydrocarbons.

Water gas is a mixture of_____

One of the catalyst mixture used in fischer-tropsch process is_____

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This water gas when hydrogenated and passed over a catalyst at _____ under 1-10 atm pressure yields crude oil

nCO + mH₂ → _____ +H₂O

Another parameter for knowing good gasoline is through ____

_____ is given an octane number of 100

_____ is given an octane number of 100

_____ is assign octane number 0

Good quality motor fuels used in modern automobiles have octane number in the range of _____

It has been observed that branching of the hydrocarbon chain increases octane number. True or false?

Octane number increase as the chain length increase. True or false?

Unsaturated hydrocarbons have lower octane number than the saturated ones. True or false?

Cycloalkanes have higher octane number than the corresponding acyclic alkanes. True or false?

The use of _____ as octane number enhancer is being curtailed for environmental reasons.

In _____, fuel having a lower octane number is much more useful than those having a higher octane number.

Quality of diesel fuel is expressed in terms of a parameter called_____.

_____is given a cetane number 100

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Good quality diesel fuel required for modern diesel engine has cetane number greater than _____

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Boiling point of a covalent substance depends upon the_____ forces.

It has been found that molecules with odd number of carbon atoms have lower melting point than those with an even number of carbon atoms. True or false?

The _____ in a carbon chain with an odd number of carbon atoms lies on the same side whereas those in a carbon chain with an even number lie on the opposite side.

All alkanes are lighter than water. True or false?

Alkanes are soluble in polar solvents but insoluble in nonpolar solvents. True or false?

UV spectroscopy is of important in the characteristic of alkanes. True or false?

The joining of the two alkyl groups from two molecules of alkyl halide with the loss of halogen occur in which method of preparation of alkanes. _____

Preparation of alkanes from carboxylic acid is achieved by _____

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Alkanes or cycloalkanes can be prepared by ____ using platinum and palladium as a catalyst.

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In Sabatier senderen’s reaction method, the hydrogenation of alkanes takes place in the presence of _____ catalyst.

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Alkylmagnesium halide is also called _____

In decarboxylation of carboxylic acid, the alkanes produced contain one carbon atom less than the original acid. True or false.

Cyclopentanone is prepared from which salt_____.

When an alkene reacts with borane, addition to the carbon-carbon double bond takes place to yield an _____

_____ can also be carried out by reacting ethyne and Grignard reagent, followed by the action of an alkyl halide.

Bond length decreases with increase in size of bonded atoms.

The chemical reactivity of an organic compound totally depends on bond length.

The structure which shows how various atoms are connected to each other is called_____.

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The formula CH₃CH₃represents the _____for ethane.

Give the name of this compound. CH₃(CH₂)₆CH₃_.

The symbol SP³ hybrid simply means

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What type of hybridization is peculiar to ethylene?

Why is the bonds formed in SP³ hybrid orbitals stronger and stable compare to the bonds formed by using pure atomic orbitals?

What is the bond angle between two SP³ hybrid orbitals?

What is the name of the bond formed between carbon and hydrogen (C–H)?

What is the measure of the length of C–C bond?

The bond energy C–C is equivalent to_____.

Which of these statements is correct?

Indicate the corresponding bond resulting from SP³, SP² and SP hybridization.

Choose the correct option that best indicate the hybrid orbitals type, bond length and bond angle for methane.

When a compound has carbon-nitrogen double bond it is called_____.

When a compound has carbon-nitrogen triple bond it is called_____.

Depending on the number of alkyl group attached to the nitrogen atom, the amines are classified as_____.

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An alcohol in which the oxygen atom is replaced by a sulphur atom is called ____.

An aromatic compound which contained side chain hydroxyl group is called_____.

The earliest nomenclature of organic compounds were based on _____.

What are isomers?

n-butane means___.

Iso-butane means_____.

The number of possible isomers of an alkane increases with increase in number of carbon atoms. True or false.

A member of a homologue series must poses a similar structure but differ in the ___ repeating unit.

‘undecane’ is a straight chain alkane containing how many carbon atoms?

In _____ reaction, a conjugated diene is treated with an unsaturated compound called dienophile to yield a cyclic system.

The terminal alkynes on hydroboration give _____.

Reactions that lead to the attachment of alkyl group to a molecular fragment are called _____.

_____ involves elimination of the halogen atom together with a hydrogen atom from an adjacent carbon atom.

Alkyl halides are converted into alkenes by _____, by treating with a strong base.

Rapid decolourization of bromine solution serves as a test for the presence of the _____ in a compound.

When alkene reacts with borane, addition to the C=C takes place to yield organoborane a compound with a carbon-boron bond, the reaction is known as _____.

In _____ compounds, the molecules are formed by the sharing of electron pairs between the constituent atoms.

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higher

C5-C7

fission
carbon monoxide and hydrogen

cobalt

Magnesia

470-870 K
Mixture of hydrocarbons

ocatane number
2,2,4-trimethylpentane
Iso-octane
n-heptane

87-95

true

False

False

true

Tetraethyllead

diesel engine

Cetane number
Hexadecane
C16H34

45
Forty five

* intermolecular*

true

Terminal carbon atoms
true

false

false

wurtz
Kolbe’s electrolytic

hydrogenation of unsaturated hydrocarbons

Nikel
Ni

grignard reagent

true
barium adipate

* organoborane*

Alkylation

False	1
Partially true	1
Lewis structure	1
spatial structure	1
Condense structure	1
Octane	1
One S and three P orbitals interaction	1
25 % S and 75 % P orbitals	1
SP² hybridization	1

In SP³ hybrid orbital, the electron density is concentrated in one direction leading to greater overlap.	1
109.5^o	1
Sigma bond	1
154 pm	1
348 Kjmo⁻¹	1
All of the options	1
Single bond, double bond and triple bond.	1
SP², 134 pm and 120^o	1
Imine	1
Nitrile	1
Primary, secondary and tertiary amines	1
First degree, second degree and third degree amin	1
Thiol	1
Phenol	1
Their origin or properties	1
Are compounds having the same molecular formula but different structural presentation	1
Straight chain butane	1
Branched butane	1
True	1
–CH₂–	1
11	1
Diels-Alder reaction	1
aldehydes	1
Alkylation reaction	1

Dehydrohalogenation	1
dehydrogenation	1
C=C	1
hydroboration	1
covalent	1

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