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# Chemistry

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Question 72

Not yet answered

Flag question

Which of the following tells the exact number of atoms of different elements present in a molecule?

a. Molecular mass

b. Empirical formula

c. Formula mass

d. Molecular formula

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# Chemistry

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Question 14

Not yet answered

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The masses of different substances, liberated or deposited by the same amount of electricity, are proportional to their equivalent masses. This statement is _____.

a. Faraday's first law of electrolysis

b. the law of definite proportion

c. Faraday's second law of electrolysis

d. the law of multiple proportion

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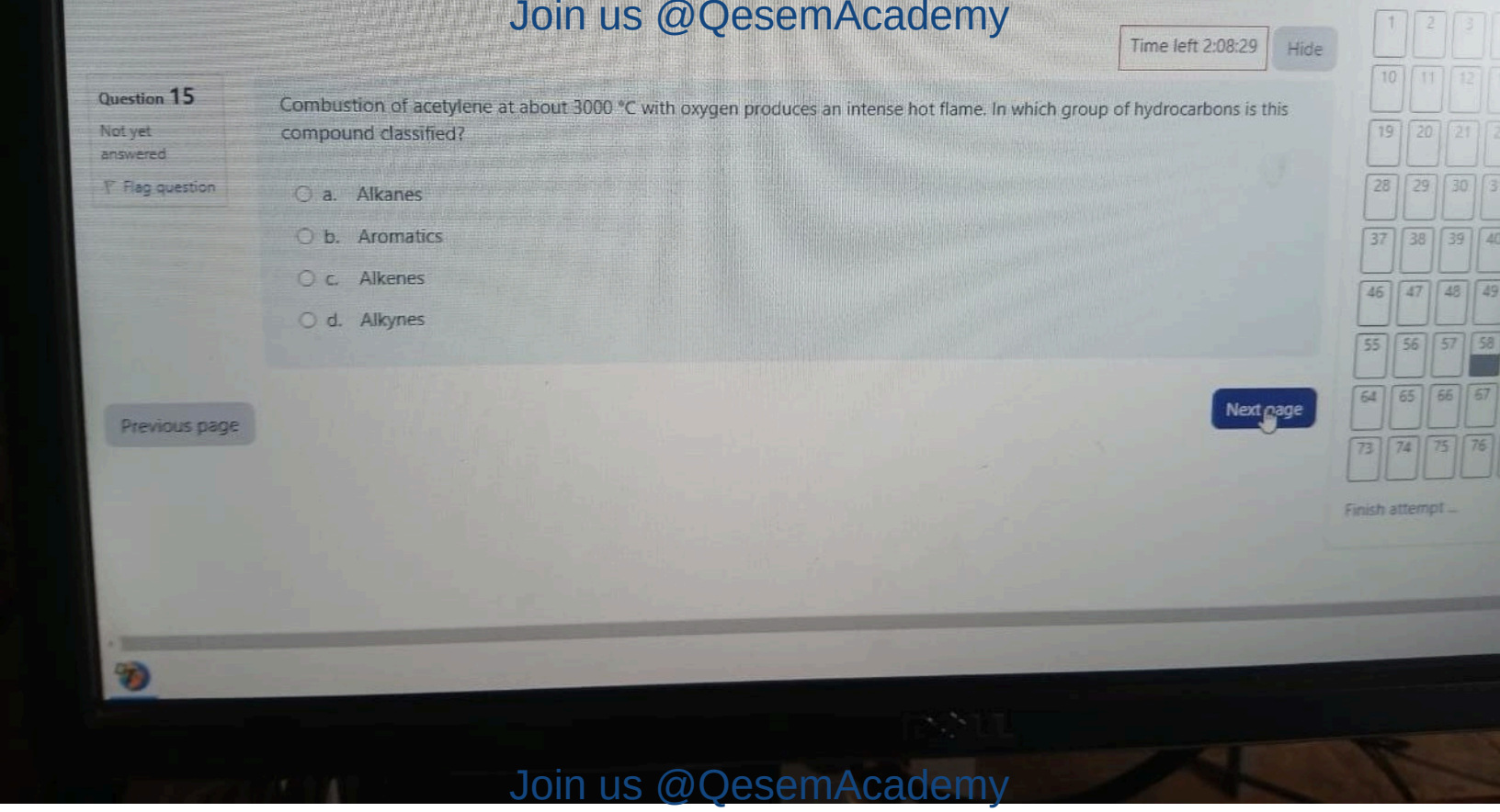


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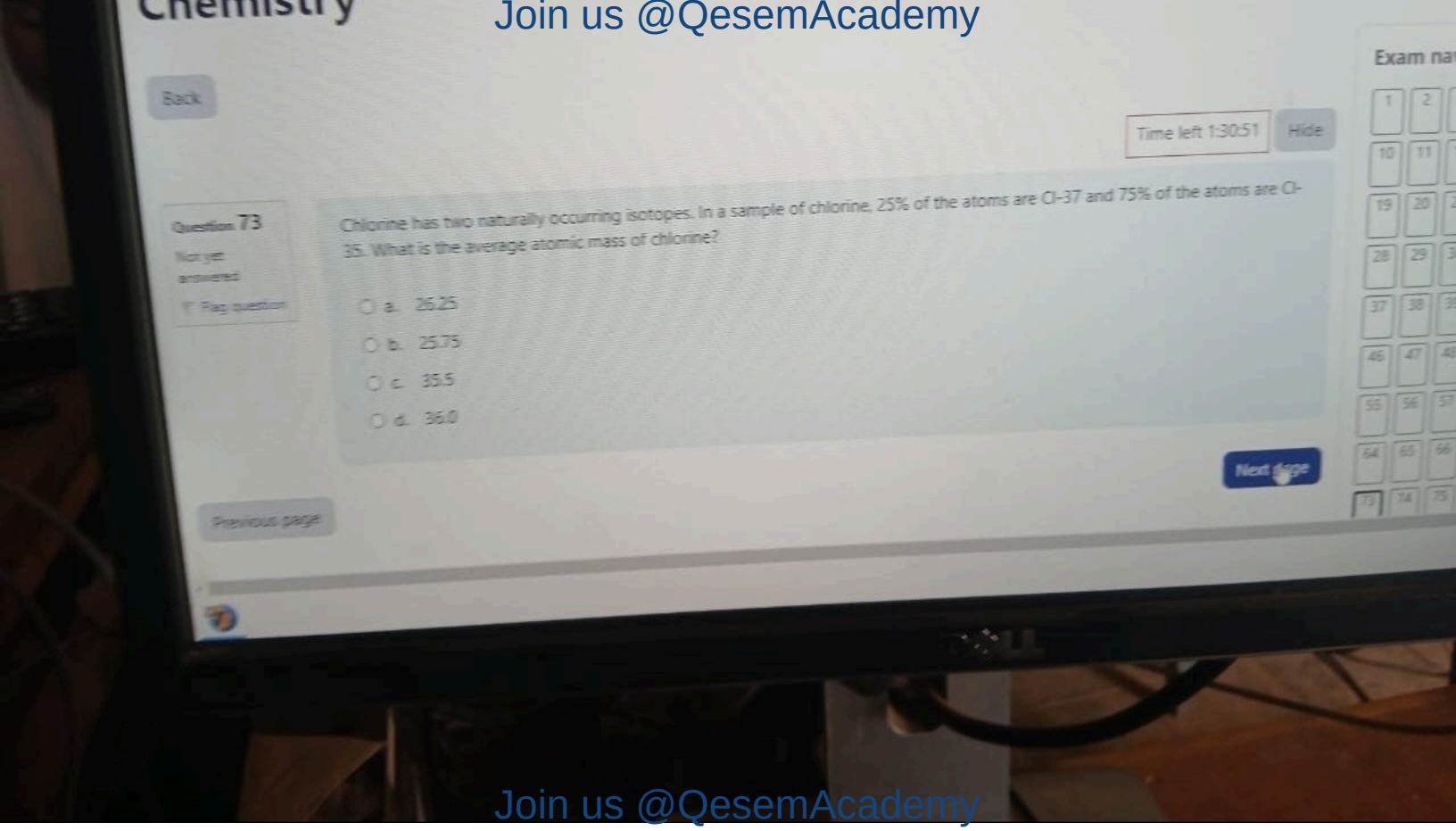


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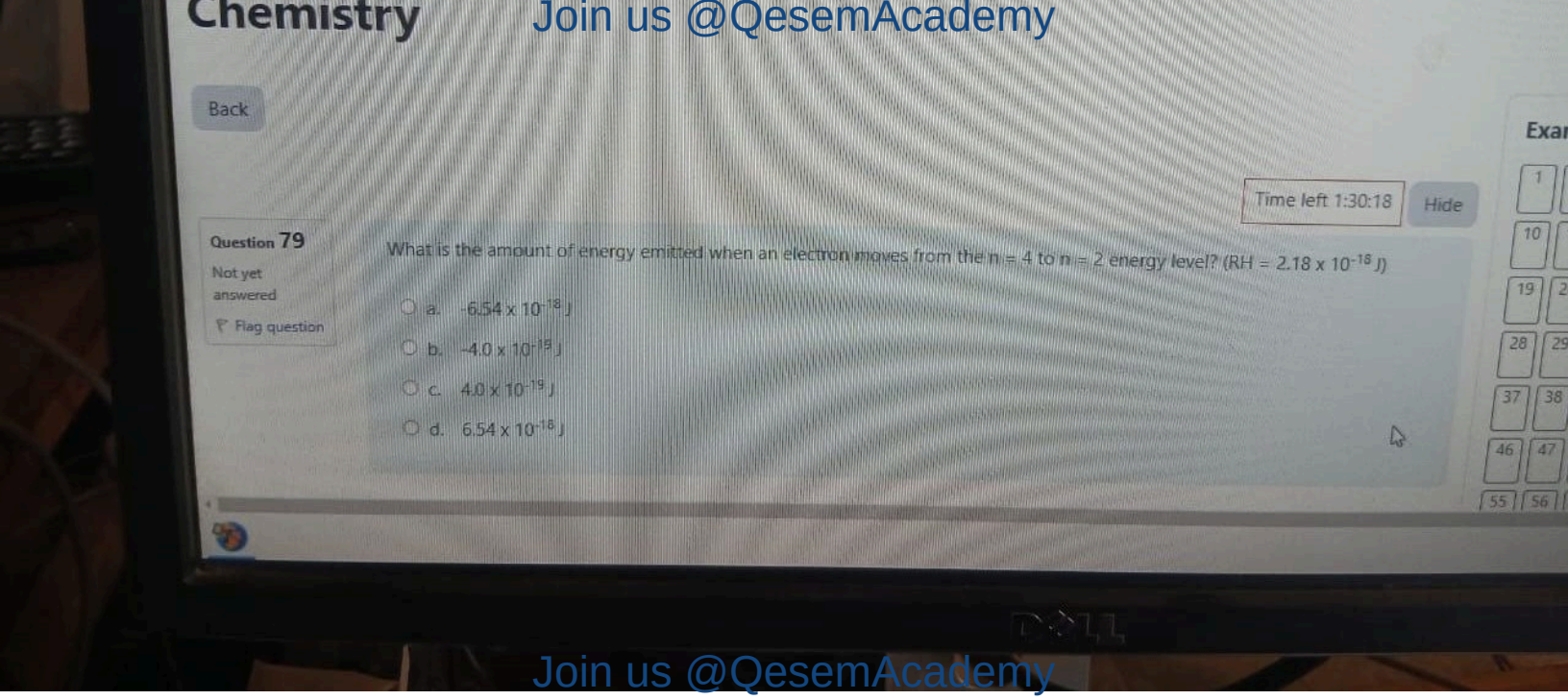
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Question 79
Not yet answered
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What is the amount of energy emitted when an electron moves from the n = 4 to n = 2 energy level? (RH = 2.18 x 10^-18 J)
a. -6.54 x 10^-18 J
b. -4.0 x 10^-19 J
c. 4.0 x 10^-19 J
d. 6.54 x 10^-18 J
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# Chemistry

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Question 65

Not yet answered

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Which one of the following does **NOT** explain the chemical properties of acidic oxides? Acidic oxides react with

a. acids to form salt and water.

b. basic oxides to form salts.

c. water to form acids.

d. base to form salt and water.

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# Chemistry

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**Question 70**

Not yet answered

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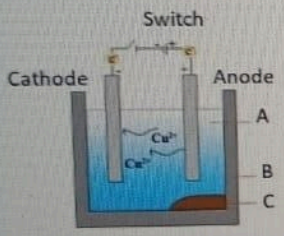
Consider the following diagram used for the electrolytic refining of copper:




What is the name of species indicated by letter 'C'?

a. Copper sulphate solution

b. Strip of pure copper

c. Cathode mud

d. Anode mud

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**Question 54**

Not yet answered

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Consider the following equilibrium reaction of a certain weak acid, *RCOOH*:

$RCOOH\ (aq) + H_2O\ (l) \rightleftharpoons RCOO^-\ (aq) + H_3O^+\ (aq)$

Which of the following is the **CORRECT** expression of the percent ionization of *RCOOH*?

a. $\frac{\text{Equilbrium concentration of } H_3O^+}{\text{Initial concentration of the acid}} \times 100$

b. $\frac{\text{Initial concentration of the acid}}{\text{Equilbrium concentration of the acid}} \times 100$

c. $\frac{\text{Initial concentration of } RCOO^-}{\text{Initial concentration of the acid}} \times 100$

d. $\frac{\text{Equilbrium concentration of the acid}}{\text{Initial concentration of the acid}} \times 100$

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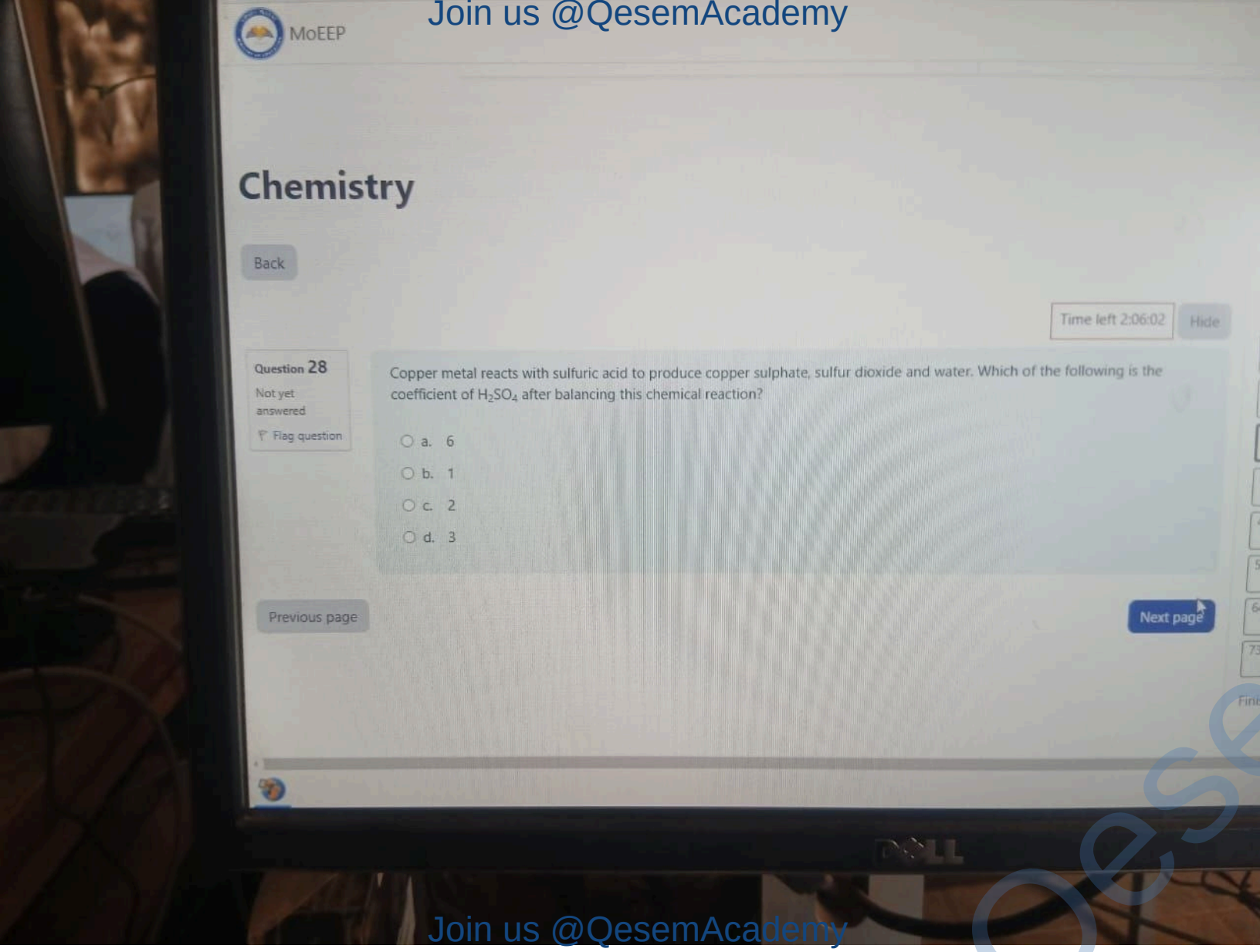


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# Chemistry

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Question 53

Not yet answered

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Consider the reaction given below:

$AB_2C + C \rightarrow AB_2 + C_2$

Which of the following successful collisions are effective in the formation of the products?

a. The collision that results in the A and C atoms are directly facing each other.

b. The collision that results in the C atoms are directly facing each other.

c. The collision that results in the B and C atoms are directly facing each other.

d. The collision that results in the A and B atoms are directly facing each other.

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# Chemistry

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Question 68

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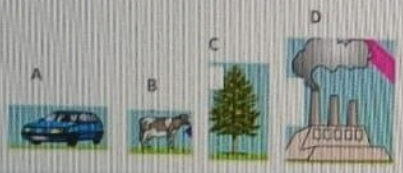
Consider the following diagram.




Which of the above does **NOT** contribute to global warming?

- a. B
- b. A
- c. D
- d. C

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Question 21

Not yet answered

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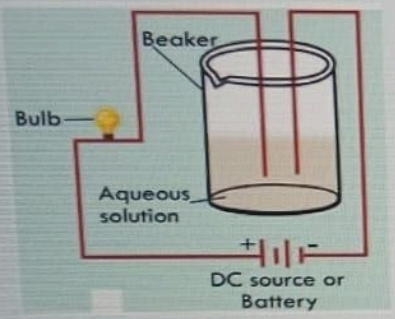
You are supposed to test the conductivity of the substances using a conductivity apparatus as shown in the figure below:




Which of the following aqueous solutions does **NOT** cause the bulb to glow when the switch turns on?

a. Sodium hydroxide solution

b. Copper sulphate solution

c. Table salt solution

d. Sugar solution

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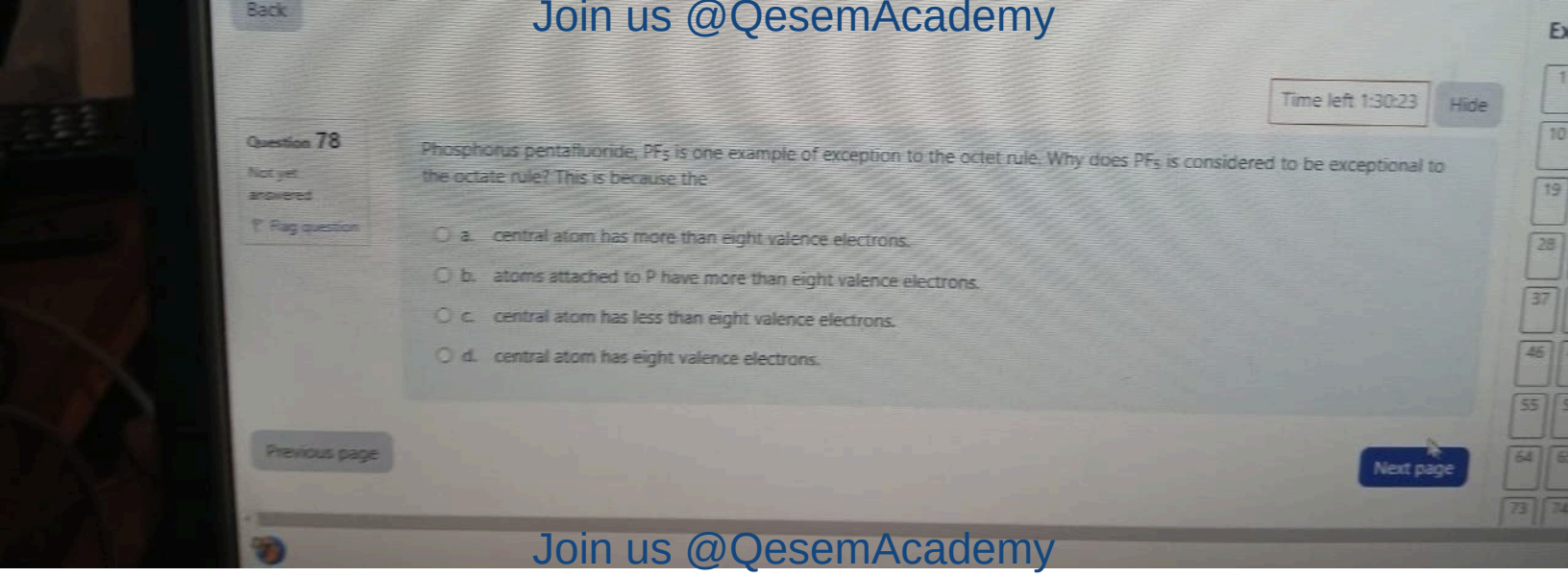


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# Chemistry

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**Question 57**

Not yet answered

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Nitric acid is a colorless fuming liquid with a boiling point of 83 °C. However, it turns yellow as it stays for long time. Which of the following accounts for the observed property of nitric acid?

- a. $NO_2$
- b. NO
- c. $NO_2^-$
- d. $NO_3^-$

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# Chemistry

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**Question 59**

Not yet answered

Flag question

A student performs an experiment to distinguish different oxides by doing the following experiment. He/she put 100 mL KI solution in to four different test tubes. He/she acidified each of these solutions with 10 drops of dilute $H_2SO_4$ solution. He/she added a spatula full of CaO in the first, $Al_2O_3$ in the second, $Na_2O_2$ in the third and $P_4O_{10}$ in the fourth test tubes. Finally, he/she added 10 mL starch solution to each of the four test tubes. In which test tube will he/she observe a color change? In the test tube that contains

a. Test tube 2

b. Test tube 3

c. Test tube 4

d. Test tube 1

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# Chemistry

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**Question 12**

Not yet answered

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A cell containing $AlCl_3$ as an electrolyte, is placed in series with another cell containing $AgNO_3$ as an electrolyte. What mass of Ag is deposited, when 18 g of Al is deposited at the cathode electrode? ($M_{Ag}$ = 108 g/mol and $M_{Al}$ = 27 g/mol)

- a. 36 g
- b. 24 g
- c. 216 g
- d. 72 g

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# Chemistry

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**Question 42**

Not yet answered

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Which of the following reaction represents the chemical equation for the reaction that takes place between aluminum metal and solution of hydrochloric acid?

a. $2AlCl_3(aq) + 3H_2(g) \rightarrow 2Al(s) + 6HCl(aq)$

b. $2Al(s) + 6HCl(aq) \rightarrow 2AlCl_3(aq) + 3H_2(g)$

c. $2Al + 6HCl(aq) \rightarrow 2AlCl_3(aq)$

d. $Al + HCl \rightarrow AlCl_3$

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# Chemistry

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Question 49

Not yet answered

Flag question

At 0 °C and 1.0 atm, the solubility of $O_2$ gas in water is 48.4 mole per liter. What is the molar solubility of $O_2$ in a saturated water solution at a partial pressure of 0.20 atm?

a. $1.20 \times 10^{-3}$ M

b. 9. 68 M

c. $4.10 \times 10^{-3}$ M

d. 2.44 M

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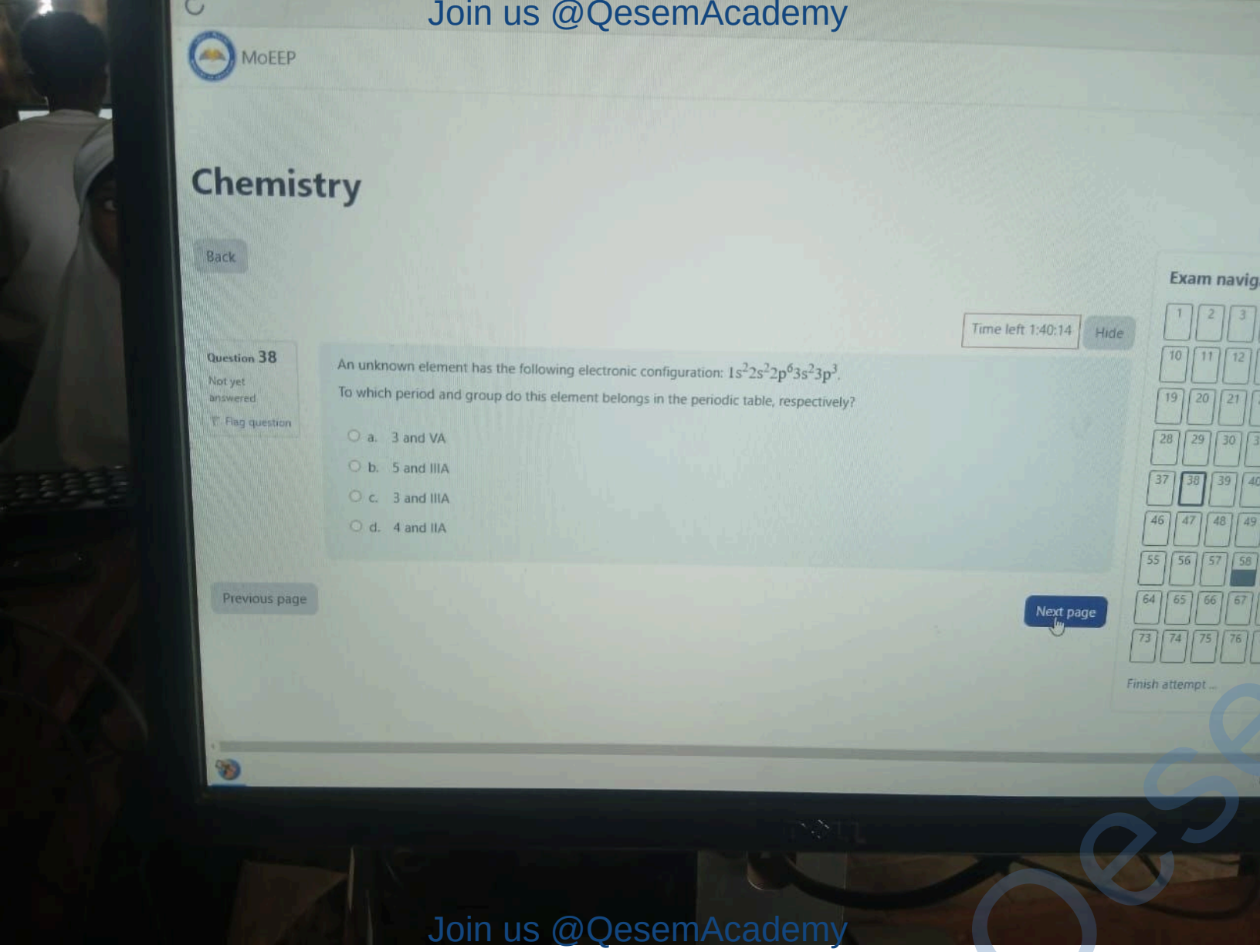


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# Chemistry

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**Question 39**

Not yet answered

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What is the concentration of hydrogen ion in a solution of 0.1 M nicotinic acid, $HC_6H_4NO_2$, at 25 °C? ($K_a = 1.4 \times 10^{-5}$)

- a. $1.2 \times 10^{-5}$ M
- b. $1.2 \times 10^{-3}$ M
- c. $1.4 \times 10^{-5}$ M
- d. $1.4 \times 10^{-6}$ M

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