ADDIS HIWOT SCHOOL, DIRE DAWA

CHEMISTRY TEST ONE FOR GRADE 7

NAME_____GRADE_____SECTION_____ROLL.NO._____

I SAY TRUE OR FALSE

1. The atomic numbers of elements were assigned by Mosley.

2. Mendeleev periodic classified elements in the order of increasing atomic number.

3. Elements in a given period have different properties because they different valence electrons.

4. Elements in the same family have the same number of shells.

5. The period number of an element corresponds to the number of shells in its atom.

II. CHOOSE THE CORRECT ANSWER FROM A GIVEN ALTERNATIVES

1. The atomic masses of the two stable isotopes of boron, 10B(20%) and $_{11}$B(80%) are 10 and 11 respectively. What is the average atomic mass of boron? a/ 11amu b/ 10amu c/ 10.8amu d/ 12amu

2. Atoms of the same element with the same number of protons but different in number of neutrons are; a/ atomic mass b/ mass number c/ isotopes d/ atomic number

3. Which one of the following scientists arranged the elements based on their atomic masses?

a/ Dobereiner's b/ Newlands c/ Mendeleev d/ all are correct

4.An element with atomic number 8 is found in---------- a/ group IIA and period 8 b/ group VIA and period 2 c/ group IIA and period 3 d/ group VIIIA and period1

5. All elements in the same period have the same----------a/ atomic number b/ physical properties c/ number of shell d/ number of valence electrons

6 .How many elements are there in the fourth period of periodic table? a/ 8 b/ 18 c/ 4 d/ 32

7. The common name of group VIIIA element is---------- a/ noble gas b/ chalcogens c/ halogens d/ alkaline earth metals

8. ----------is an atom or group of atoms that is positively or negatively charge?

a/ molecule b/ atom c/ Iron d/ Ion

9. The actual mass of neutron is----------

a/ 1amu b/ 9.11×10^{-28}g c/ 1.673×10^{-24} g d/ 1.675×10^{-24}g

10. The relative mass of electron is---------- a/ ≈ 0 b/9.11×10^{-28}g c/1amu d/ 1.675×10^{-24}g

III.WORK OUT

1. Balance the following chemical equations (you can use inspection, Algebraic or LCM method).

a. $N_2 + H_2 \rightarrow NH_3$

b. $Al + Cl_2 \rightarrow AlCl_3$

c. $Ca + HCl \rightarrow CaCl_2 + H_2$

d. $2Al + Fe_2O_3 \rightarrow Al_2O_3 + 2Fe$

e. $Cu + HNO_3 \rightarrow Cu(NO3)_2 + NO + H_2O$