F1 ENDYEAR ASSIGNMENT

CHEMISTRY

10+ HOLIDAY ASSESMENT PAPERS

Dear Students, Make Maximum Utilization of limited time Available to Attempt and Work on these Year-Round Examination Trials.

A Comprehensive Evaluation Analysis of Term 1-3 Questions.

SERIES 1

For Marking Schemes Call/WhatsApp 0746 222 000

MWALIMU AGENCY

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 1

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

1. **(a)** State one disadvantage of using flower extracts as acid – base indicators. **(1mk)**

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(b) Name the indicator that can be used in the laboratory to tell the pH of lemon juice. **(1mk)**

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(c) Differentiate between strong and weak acids. **(1mk)**

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2. **(a)** What are isotopes? **(1mk)**

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(b) Determine the number of neutrons in $^{18}_{8}O$ **(1mk)**

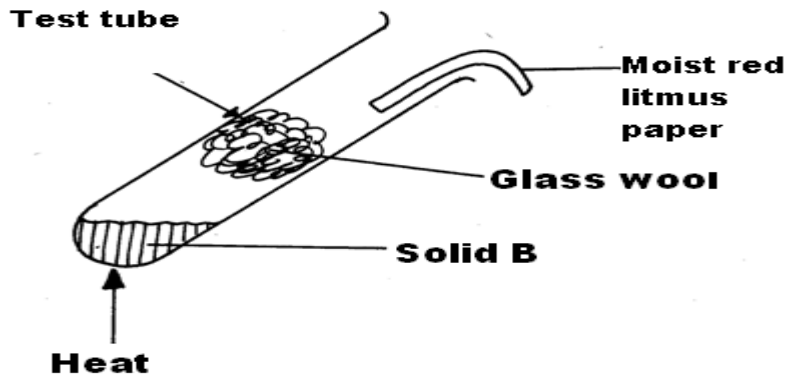
3. When magnesium is burnt in air it reacts with both oxygen and nitrogen gas giving a white ash. Write two equations for the reactions that take place. **(1mk)**

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4. A solution contains 29.1g per litre of aluminium Sulphate. Calculate the number of Sulphate ions in 350cm^3 of the solution.(Al = 27, S = 32, O = 16) Avogadro's constant = 6.0 x 10^{23}. **(2mks)**

5. When a solid B was heated in a test-tube, it gave off two gases. The two gases were seperates by passing them through a plug of glass wool in a test-tube as shown below.



The first gas which evolved turned moist red litmus paper to blue. Later the other gas involved turned the litmus back to red.

(a) Identify solid B **(1mk)**

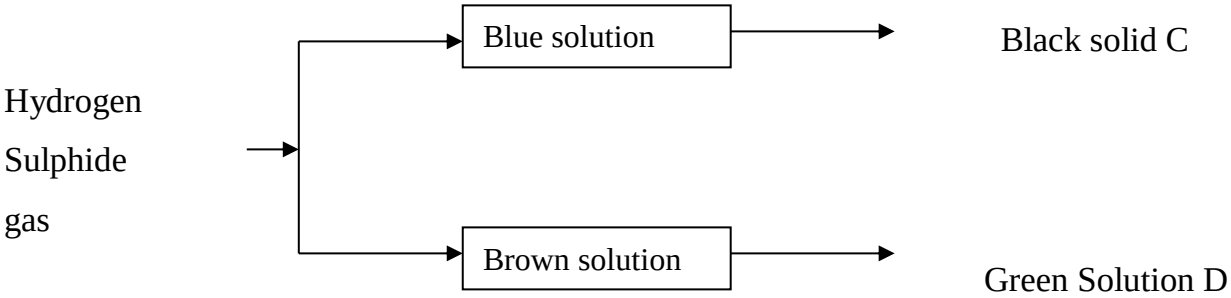
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(b) Write the equation for the reaction that take place in the test tube **(1mk)**

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6. Hydrogen Sulphide gas was bubbled into two solutions of metallic nitrates as shown in the flow diagram below



(a) Identify the black solid C **(1mk)**

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(b) Write an ionic equation for the formation of the green solution **(1mk)**

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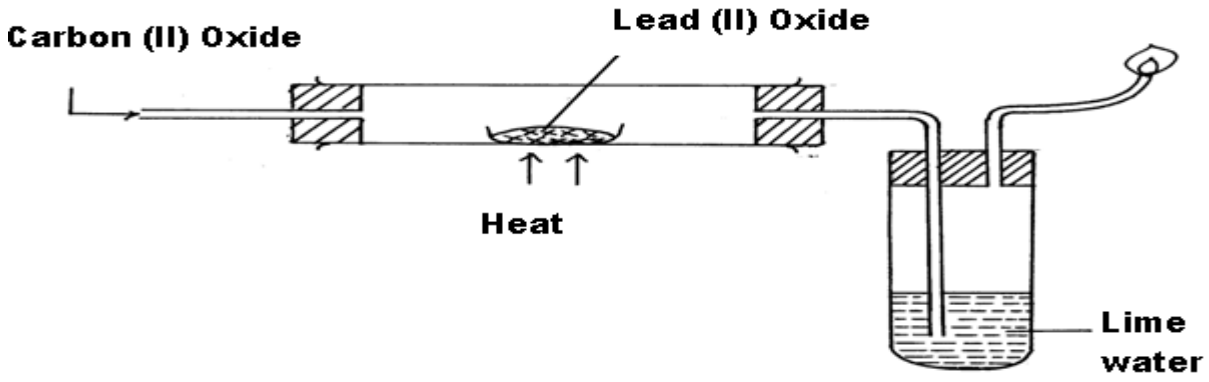
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(c) State the property of Hydrogen Sulphide shown by the formation of solution **D** **(1mk)**

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7. The apparatus shown below was used to investigate the effect of Carbon (II) Oxide on Lead (II) Oxide.



(a) State the observation made in the combustion tube during the experiment. **(2mks)**

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(b) Write the equation for the reaction that take place in test-tube E. **(1mk)**

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8.(a) Lime water is also used to soften hard water. Why is this method not preferred to soften hard water. **(1mk)**

……

(b) Name a compound that causes temporary hardness of water. **(1mk)**

……

(c) State one disadvantage of using hard water for domestic purpose. **(1mk)**

……

9. **(a)** Name the compound below

$CH_3CH_2CH = CH_2$

……

……

(b) Draw and name other isomers of the compound in (a) above. **(2mks)**

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10. A white solid dissolve in water to form a colourless solution. The colourless solution forms a white precipitate with Ammonia solution but dissolve in excess alkali. The colourless solution forms a white precipitate with Lead (II) Nitrate solution. The white Precipitate dissolve on warming to form a colourless solution.

(a) Write the chemical formulae for the ion formed when the colourless solution react with excess ammonia solution. **(1mk)**

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(b) Write the name of the ion present in the white solid. **(1mk)**

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(c) What is an alkali **(1mk)**

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11. The solubility of Potassium Manganate (VII) at 20^0C is 13g per 100g of water and at 90^0C is 60g per 100g of water.

(a) Determine the mass of Potassium (VII) Manganate present in 80g of saturated solution at 90^0C. **(1mk)**

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(b) Calculate the mass of Potassium (VII) Manganate that would crystallize out if the solution in (a) were cooled to 20^0C. **(2mks)**

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12. (a) Explain why concentrated Sulphuric (VI) acid is a poor electrolyte and has no effect on blue litmus paper whereas 2M Sulphuric (VI) acid conducts electricity and changes blue litmus paper red. **(1mk)**

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(b) What is a binary electrolyte. **(1mk)**

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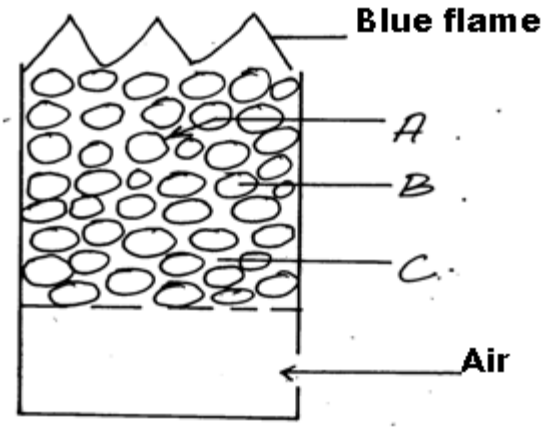
13. An element F has a relative atomic mass of 88. When a current of 0.5 amperes was persed through the fused chloride of F for 20 minutes and 20 seconds, 0.278g of F were deposited at the cathode. Determine the charge on ion of F (1Faraday = 96500C). **(2mks)**

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14. Below is a cross- section of a charcoal burner



(a) Charcoal is a form of impure carbon. Name any other two allotropes of carbon. **(1mk)**

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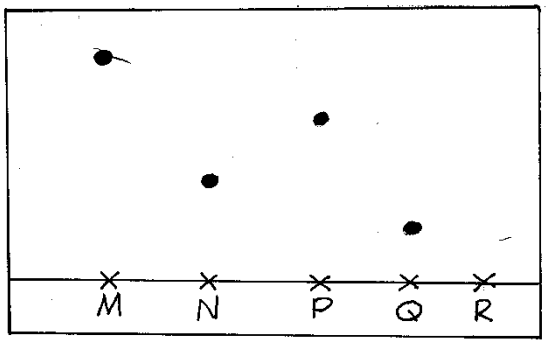
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(b) Write an equation for the reaction taking place at the part marked B. **(1mk)**

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15. The diagram below represents paper chromatogram of four types of sugar.



(a) Identify the most stable sugar. **(1mk)**

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(b)On the diagram, show the chromatogram of R. Given that R contains all the other sugar**s (2mks)**

16. The empirical formula of a hydrocarbon is C_2H_3. The hydrocarbon has a relative molecular Mass of 54. (H = 1, C = 12)

(a) Determine the molecular formula of the hydrocarbon. **(1mk)**

(b) Draw the structural formulae of the hydrocarbon in (a) **(1mk)**

(c) To which homologous series does the hydrocarbon in (b) above belong? **(1mk)**

17. State three factors that increase the rate of reaction for the following reaction. **(3mks)**

$CaCO_{3(s)} + 2HCl_{(aq)} \xrightarrow{\quad} CaCl_{2(aq)} + CO_{2(g)} + H_2O_{(aq)}$

18. In terms of structure and bonding explain the following

(a) Melting point of Magnessium is higher than that of Sodium. **(3mks)**

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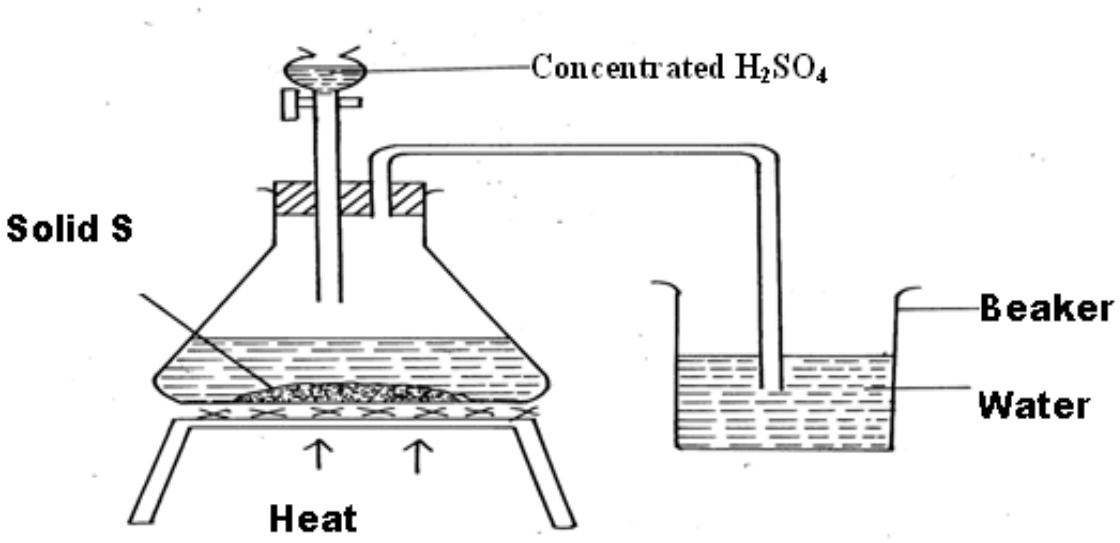
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(b) Melting point of Chlorine is lower than that of Iodine. **(1mk)**

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19. The set-up below was used to prepare a solution of hydrogen chloride gas



(a) Identify solid S **(1mk)**

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(b) Identify one mistake in the set up. **(1mk)**

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(c) Write an equation for the reaction taking place in the flask. **(1mk)**

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20. (a) State gay Lussac's law **(1mk)**

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(b) $10cm^3$ of a gaseous hydrocarbon, CxHy required 30cm3 of oxygen for complete combustion. If steam and $20cm^3$ of Carbon (IV) oxide were produced, what is the value of x in CxHy. **(2mks)**

21. Starting with lead metal, describe how a dry sample of Lead(II) chloride can be prepared in the laboratory. **(3mks)**

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22. An ion T^{2-} has an electronic arrangement of 2.8

(a) What is the atomic number of the element. **(1mk)**

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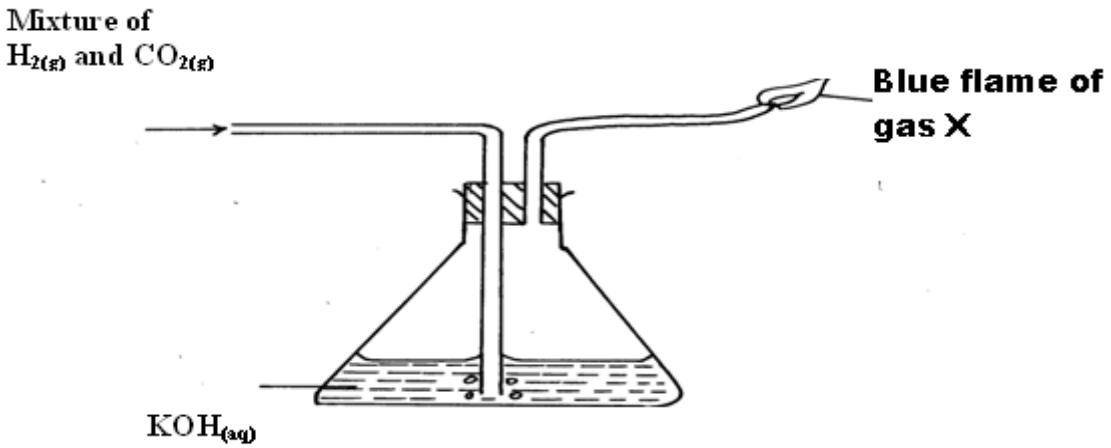
(b) To which group and period does the element belong to;

Group_____ **(1mk)**

Period_____ **(1mk)**

23. Using dot (.) and cross (x) diagram show the type of bond present in hydrogen ion, H_3O^+
(H = 1, O = 8) **(2mks)**

24. A mixture of Hydrogen gas and Carbon (IV) oxide are passed through Potassium hydroxide solution as shown below.



(a) State the observation made in the conical flask. **(1mk)**

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(b) Write the equation for the reaction that takes place in:-

(i) the conical flask **(1mk)**

……

(ii) the burning of gas x **(1mk)**

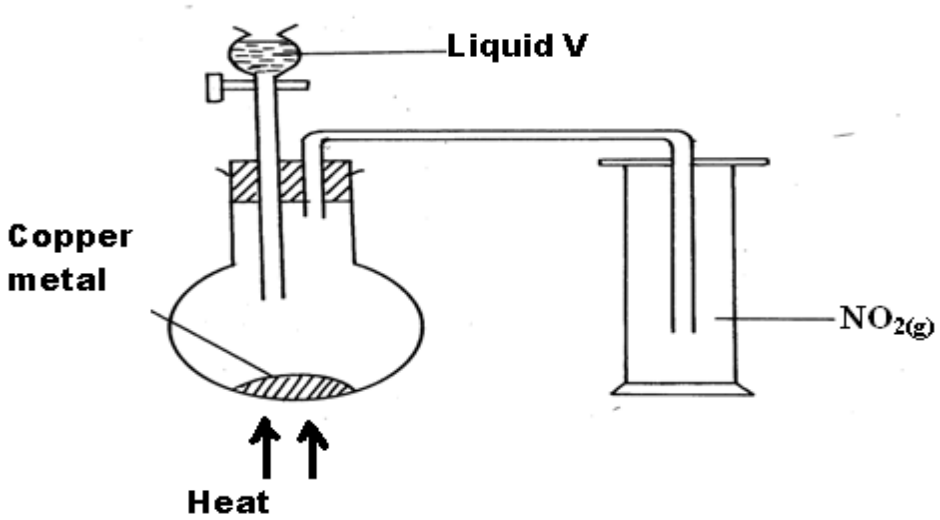
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25. $20cm^3$ of a solution containing 4g per litre of Sodium hydroxide was neutralized by 8cm3 of dilute Sulphuric (VI) acid. Calculate the concentration of the acid in moles per litre.
(Na = 23, O = 16, H=1) **(3mks)**

26. Given that the hydration energies of $Ca^{2+}_{(g)}$ and $Cl^{-}_{(g)}$ are -1562KJ/Mole and -364KJ/Mole respectively. The heat of solution (Δ Hsoln) for one Mole of $CaCl_2$ is -82.9 KJ/Mole. Determine the lattice energy for $CaCl_2$. **(2mks)**

27. The diagram below is used to prepare nitrogen (IV) oxide gas.



(a) Identify substance V **(1mk)**

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(b) State and explain one precaution taken when carrying out the experiment. **(2mks)**

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28. Electrode potentials for the half cells are shown below. Use the information to answer the questions that follow.

$Sn^{2+}_{(aq)} + 2e^- \rightarrow Sn_{(s)}$; $E^\theta = -0.14V$

$Cu^{2+}_{(aq)} + 2e^- \rightarrow Cu_{(s)}$; $E^\theta = -0.34V$

(a) Write the cell representation for the cell made up of two half cells. **(1mk)**

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(b) Write the cell equation for the cell reaction. **(1mk)**

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(c) Calculate the E^θ value for the cell. **(1mk)**

29. State the function of each of the following in the Solvay process of production of Sodium Carbonate.

(a) Coke **(1mk)**

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(b) Cold water on the carbonation. **(1mk)**

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(c) Ammonia a generator. **(1mk)**

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30. (a) A student in form three was given two gases C_2H_6 and C_2H_4. He added acidified Potassium Manganate (VII) to each solution. State the observations the student made. **(2mks)**

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(b) State one use of C_2H_4. **(1mk)**

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FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 2

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

QUESTIONS

1. (a) Define the term chemistry **(2mks)**

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(b) List four branches of chemistry **(4mks)**

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(c) Name three main branches of science **(3mks)**

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2. Define the following terms **(5mks)**

(a) Mixture

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(b) Compounds

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(c) Elements

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(d) Atom

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(e) Molecule

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3. State two properties of

(a) Solid **(2mks)**

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(b) Liquids **(2mks)**

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4. What is a non-conductor? **(1mk)**

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5. Define the following terms

(a) A drug **(1mk)**

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(b) Drug abuse **(1mk)**

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6. (a) Name three commonly abused drugs **(3mks)**

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(b) Mention three negative effects of prolonged use of khat. **(3mks)**

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7. Name three roles chemistry plays in the society. **(3mks)**

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8. List three careers one can join after studying chemistry **(3mks)**

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9. List three scientific skills that can be acquired when performing chemistry experiments. **(3mks)**

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10. List five laboratory safety rules **(5mks)**

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11. List three apparatus that are used to measure accurate volume **(3mks)**

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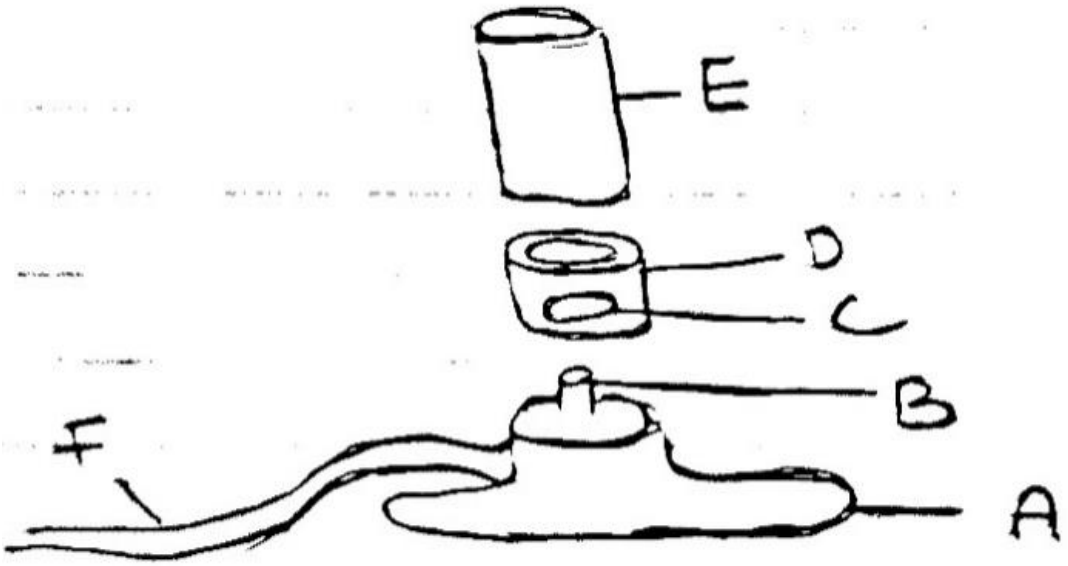
12. Draw and give the use of the following apparatus used in the chemistry **(6mks)**

(a) Thistle funnel

(b) Separating funnel

(c) Dropping funnel

13. The diagram below represents a Bunsen burner



(a) Give the name of the parts labeled A,B,C,D,E and F on the diagram **(6mks)**

A _____

B_____

C_____

D_____

E_____

F_____

(b) A Bunsen burner can produce two different types of flames under different conditions.

(i) Name the two types of flames produced by a Bunsen burner **(2mks)**

……

……

(ii) Give 3 differences between the two flames in (i) above **(3mks)**

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(iii)Draw and label the most suitable Bunsen burner flame preferred for heating in the laboratory **(3mks)**

14. Give the most suitable method of separating the following mixtures **(5mks)**

(a) Water and Ethanol

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(b) Kerosene from Water

……

(c) Colored pigments

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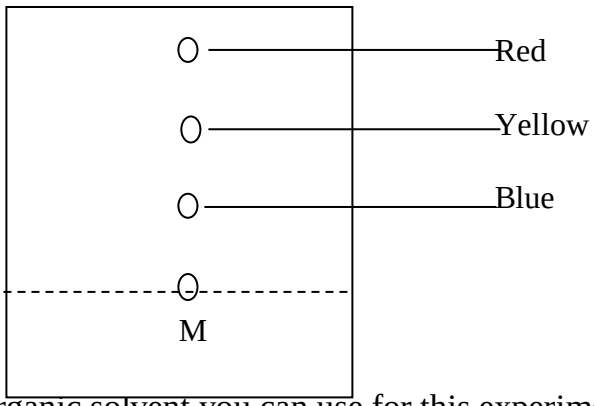
(d) Iodine from a mixture of sand and Iodine

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(e) Iron fillings from Sulphur Powder.

……

15. The chromatography below show the constituents of a flower extract using organic solvent



(i) Name a possible organic solvent you can use for this experiment **(1mk)**

……

(ii) State two properties that make the red pigment to move the furthest distance from M**.** **(2mks)**

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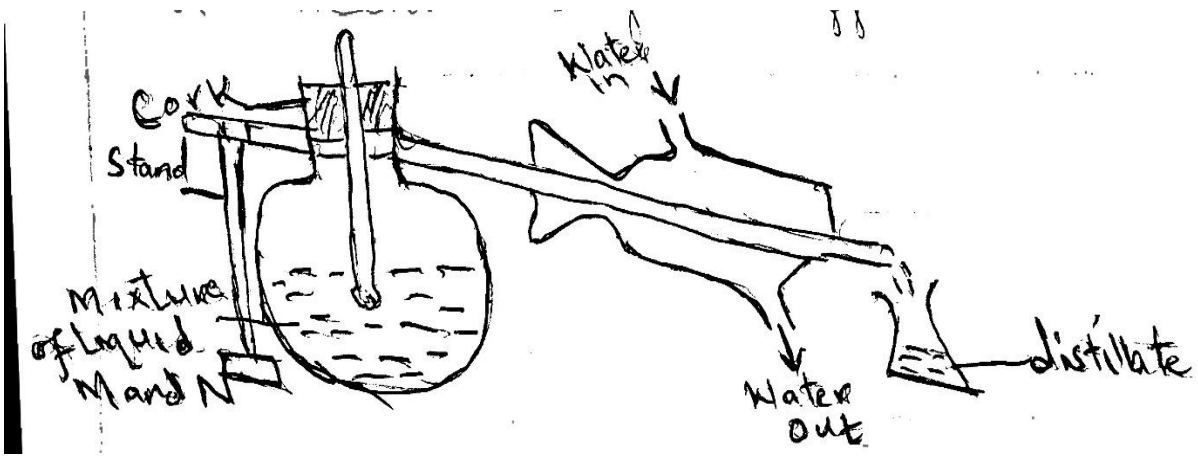
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(iii) On the diagram indicate the solvent front. **(1mk)**

16. Write down the chemical symbols of the following elements **(5mks)**

Element	Chemical Symbol
(i) Copper	……..
(ii) Zinc	……
(iii) Silver	……..
(iv) Aluminium	……
(v) Mercury	……

17. In a experiment to separate a mixture of two organic liquid M and N with boiling point of 56˚c and 118˚c respectively, a student set up the apparatus as shown



(a) Identify 3 mistakes in the setup **(3mks)**

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……

……

(b) What method would the students use to test the purity of distillate obtained **(2mks)**

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(c) What is the purpose of the thermometer **(1mk)**

……

(d) Give two applications of fractional distillation **(2mks)**

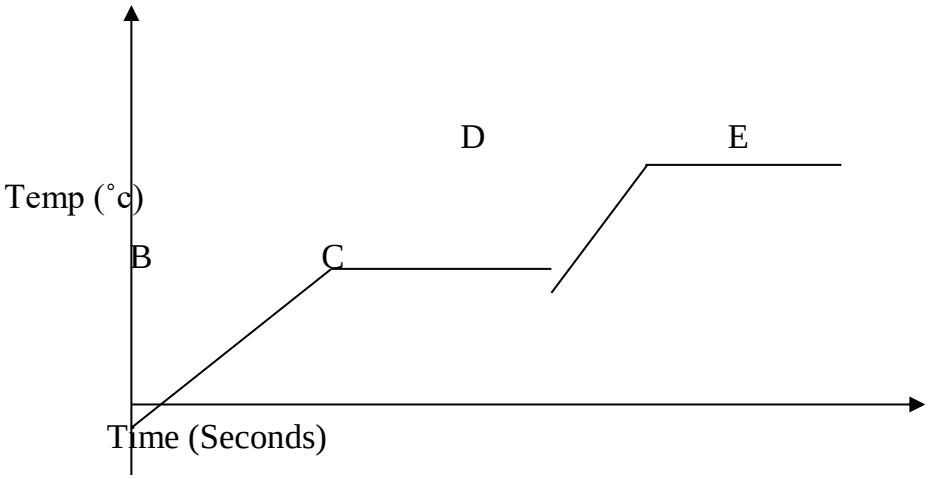
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18. Give three differences between chemical changes and physical changes **(3mks)**

Chemical change	Physical change

19. The graph shown below is heating curve of solid substance. Use it to answer questions below



(a) Using kinetic theory of matter explain what happens in the following regions **(6mks)**

(i) AB

……

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(ii) BC

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……

(iii) CD

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20. Explain the following

(a) Boiling tube is usually more suitable for boiling liquids than a test tube **(2mks)**

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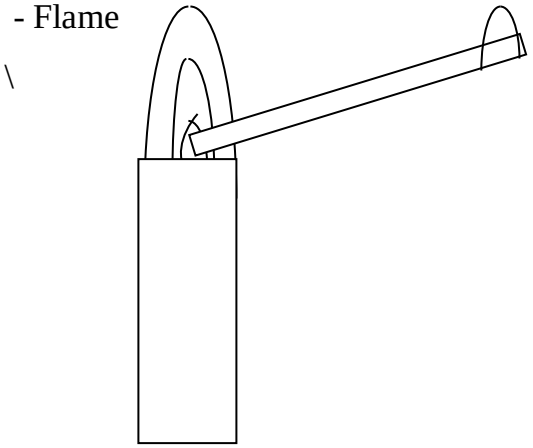
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(b) Round bottomed flask is used for fractional distillation and not a flat bottomed flask **(1mk)**

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(c) - Flame



- Noisy Bunsen burner

When glass tube is held in a noisy Bunsen flame as shown a small flame appears at the end of the glass tube. Explain **(2mks)**

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FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 3

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

QUESTIONS

1. Define the following terms **(4marks)**

a) Drug

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b) Prescription

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c) Nekesa visited a hospital and was given a syrup whose prescription was2×3.How should she take the syrup? **2marks)**

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2(a) Explain the following:

(i) It is always advisable to scoop chemical substances using a clean spatula**.** **(1mark)**

……

……

(ii) Flammable substances should always be kept away from flames in the laboratory. **(1mark)**

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(b) (i) Give ***three*** differences between luminous and non-luminous flames. **(3 marks)**

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(ii) How is the non-luminous flame produced? (1 mark)

(c) (i) Explain why solid Carbon (IV) oxide is preferred over ordinary ice for use by ice cream venders**.** **(1 mark)**

……

……

(ii) Name one piece apparatus used to measure volume of gases. **(1 mark)**

……

……

(iii) Draw a diagram of a deflagrating spoon and state its use **(2 marks)**

3. Define the following terms **(5marks)**

a) Solute.

……

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b) Solvent.

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c) Solution.

……

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d) Residue.

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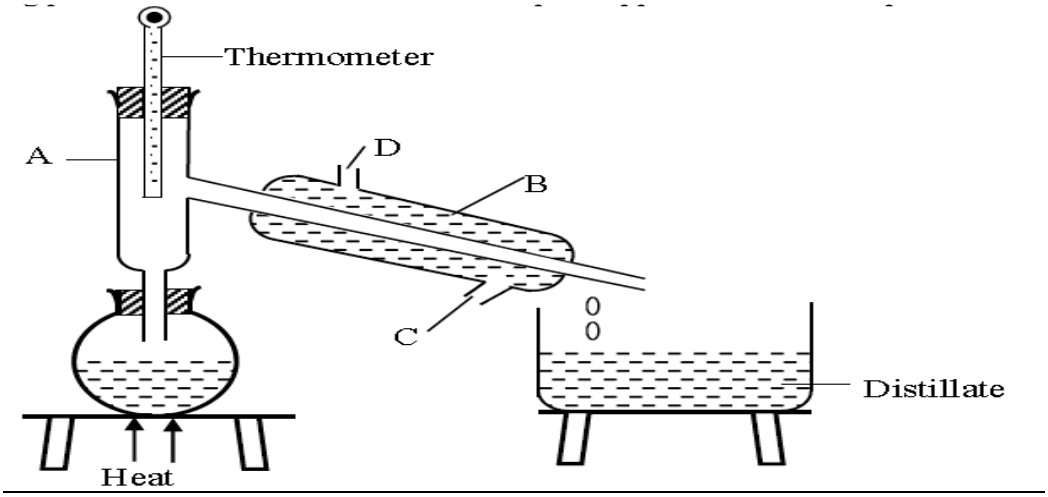
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e) Filtrate**.**

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4. (a) A form one had a mixture of ethanol and water. Ethanol has a boiling point of 78°C while water has a boiling point of 100°C. The student then set up the apparatus below to separate the mixture.



(i) Name the piece of apparatus labelled B. **(1 mark)**

……

(ii) What is the purpose of the thermometer in the set up? **(1 mark)**

……

(iii) Name the part labelled A and state its function **(2 marks)**

……

(iv) Which liquid was collected first? Explain **(2 marks)**

……

……

(v) What is the name given to the above method of separating mixtures? **(1 mark)**

……

(vi) What passes through parts C and D**?** **(2marks)**

C

D

(vi) What property of the components of the mixture makes it possible for the components to be separated by the method**?** **(1mark)**

……

(vii) State one applications of the above method of separation. **(1 marks)**

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5(a) Describe how you can extract oil from ground nuts**?** **(3marks)**

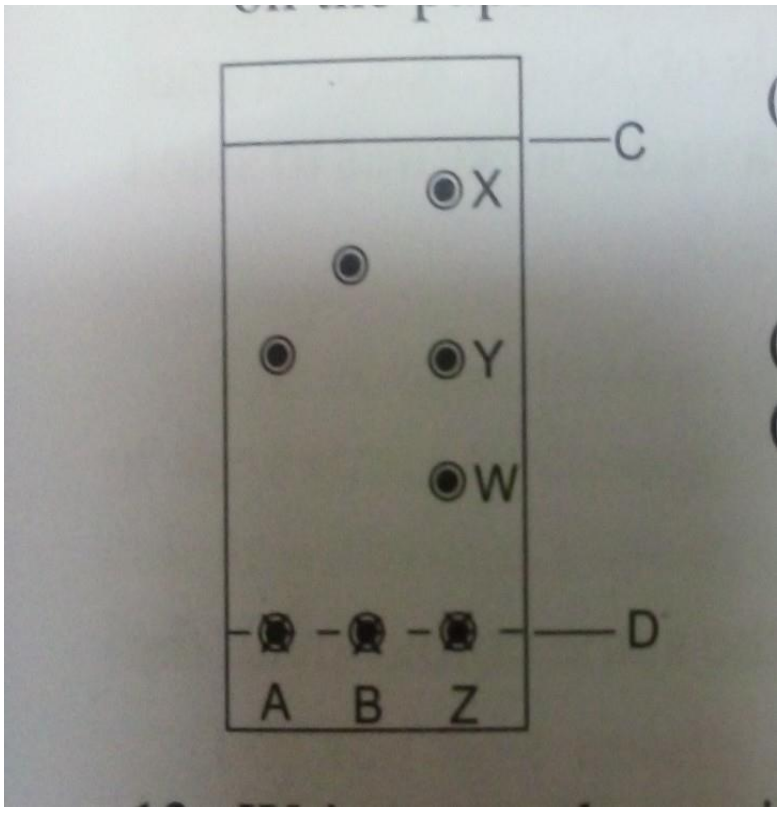
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6. Spots of pure pigments A and B, and a mixture of Z were placed on a filter paper and allowed to dry. The paper was then dipped in a solvent. The results obtained were as on the chromatogram.



a) which is the **;** **(2marks)**

i) Base line?

……

……

ii) Solvent front?

……

……

b) Which of the pure pigments was a component of Z Explain? **(2marks)**

……

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(c)(i) Name a solvent that is used in paper chromatography**.** **(1mark)**

……

(ii)Why is water not a suitable solvent in paper chromatography? **(2marks**

……

……

(iii) Give 2 applications of chromatography in our daily lives**?** **(2marks)**

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7. Explain the differences between solid, liquid and gaseous states using the theoretical model (diagram) of matter in terms of the kinetic theory **(6marks)**

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8. Define the following terms **(4marks)**

a) An atom

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b)A molecule

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c) An element

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d) A compound

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9. Complete the following table **(4marks)**

Element	Symbol
potassium	
	Na
Silver	
	Au
Iron	
	Pb
Copper	
Mercury	

10. Name the elements present in the following compounds.

a) Sodium Bromid**e** **(2marks)**

……

……

b) Magnesium nitride **(2marks)**

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……

11. Write a word equation for the reaction between:

a) Carbon and oxygen **(2marks)**

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……

b) Sodium and sulphur **(2marks)**

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12. Give three differences between a temporary chemical change and a permanent chemical change **(3marks)**

Temporary chemical change	Permanent chemical change

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 4

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

QUESTIONS

1(a). Name three frequently abused drugs**.** **(3mks)**

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(**b).**State two long term effects of drug abuse**.** **(2mks)**

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2. (i). Define Chemistry**.** **(1mk)**

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(ii) Give three importance of studying Chemistry**.** **(3mks)**

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3(a). The following are laboratory apparatus used in Chemistry. Name them and give their uses.

(i) 

name

use.

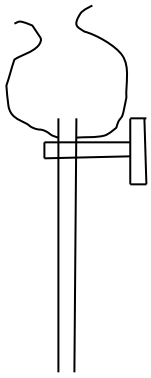
(2mks)

(ii) 

name

Use.

(2mks)

(iii) 

Name

Use. **(2mks)**

(b). Give two reasons why most laboratory apparatus are made of glass. **(2mks)**

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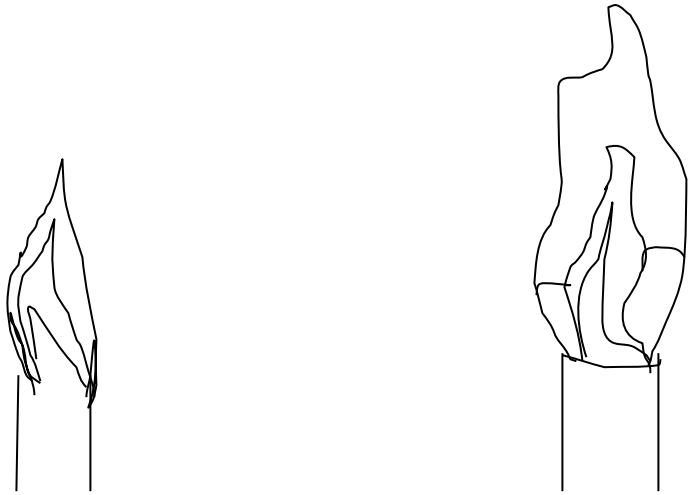
4.(i) What is a flame**?** **(1mk)**

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……

(ii). The following diagrams represent the two types of flames produced by a bunsen burner.

(a) (b)



(a). Identify the flames (a) and (b**).** **(2mks)**

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……

(b). Which type of the flames identified above is preferred for heating? Give a reason for your answer. **(2mks)**

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(c). give four differences between the flames (a) and (b) above. **(4mks)**

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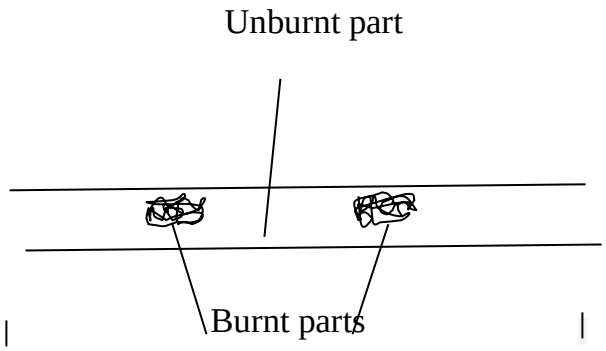
(d)(i). Give two reasons why flames (a) and (b) in 4(ii) above differ**.** **(2mks)**

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(ii).A wooden splint was slipped through a region of a particular flame of the Bunsen burner in the laboratory. The splint was burnt as shown in the diagram below.



(a). Name the type of flame the splint was slipped through**.** **(1mk)**

……

(b). Explain why the splint was burnt the way it is shown in the diagram**.** **(2mks)**

……

……

(iii). After use, the non-luminous flame should be put off or adjusted to luminous flame. Explain. **(2mks)**

……

……

(iv). Putting off flames is one of the laboratory safety rules. State four other rules. **(4mks)**

……

……

……

……

(**v).** What is a mixture? **(2mks)**

……

……

(vi). State two physical means of separating a mixture. **(2mks)**

……

……

5.Isaboke,a form one student at Buka academy accidentally mixed sulphur and iron filings.

(a). Suggest an appropriate method of separation you would advice him to use to separate the mixture. **(1mk)**

……

(b). Give a reason for the choice of your answer. **(2mk)**

……

……

(c). Describe how he would use the method named in 5(a) above ton separate the sulphur and iron filings. **(4mks)**

……

……

……

……

……

(d). Name two substances that sublime when heated**.** **(2mks)**

……

……

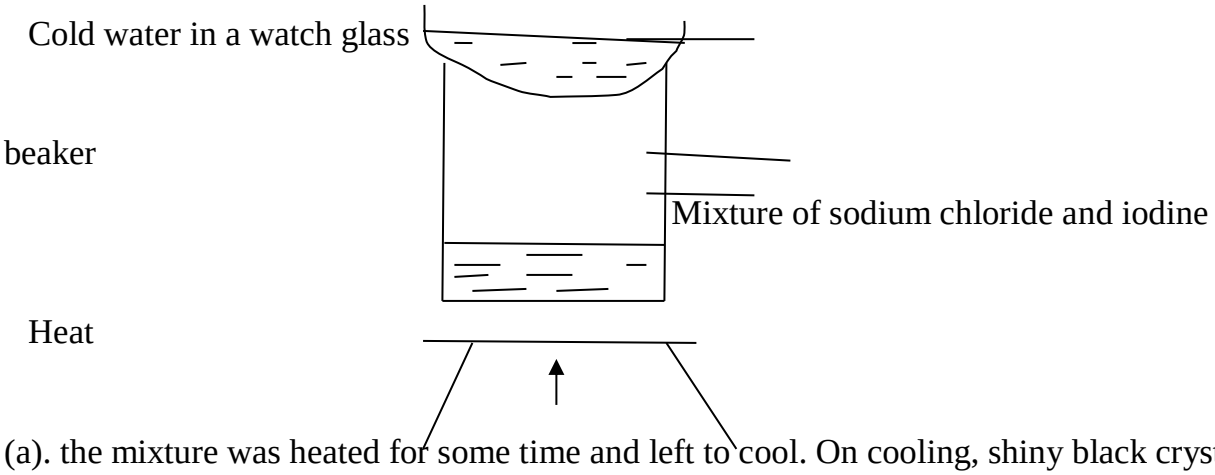
(e). Give two reasons why dry ice (solid carbon (IV) oxide) is preferred to be used in cold boxes by ice cream vendors over ordinary ice**.** **(2mks)**

……

……

……

6. Majaliwa, a form one student at Kanga High School wanted to separate and obtain iodine and sodium chloride (common table salt) from a mixture of the two. he set the experimental set up shown below.



(a). the mixture was heated for some time and left to cool. On cooling, shiny black crystals and white crystals were observed on the surface of the watch glass and in the beaker respectively. Name:

I. Shiny black crystals

II. White crystals**.** **(2mks)**

(b). What was the purpose of the cold water in the watch glass**?** **(2mks)**

……

……

(c). Explain how the shiny black crystals on the surface of the watch glass is formed. **(3mks)**

……

……

……

(d). What property of iodine makes it be collected on the watch glass as shown**?** **(1mk)**

……

……

(e). Explain why it is possible to separate a mixture of iodine and sodium chloride. **(1mk)**

……

……

7(a). Define the following terms:

1) A saturated solution**.** **(1 ½ mks)**

……

……

2) Crystallization**.** **(1 ½ mks)**

……

……

(b)(i). Describe how copper (II) sulphate crystals can be obtained from copper (II) sulphate solution. **(4mks)**

……

……

……

……

(ii). Give two industrial applications of crystallization as a method of separating soluble substances from their solutions. **(2mks)**

……

……

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 5

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

1) Explain why most laboratory apparatus are made of glass. **(2mks)**

……

……

……

2) (a) What is drug abuse? **(1mk)**

……

……

(b) Name three drugs that are commonly abused. **(3mks)**

……

……

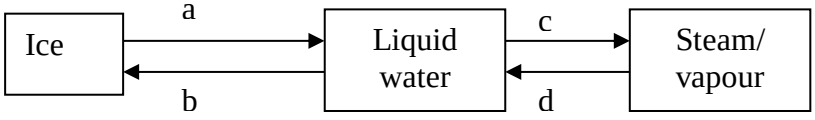
……

(c) State two ways of preventing drug abuse. **(2mks)**

……

……

3) Study the diagram below and answer the question that follows.



a) Name the process labeled a, b, c and d. **(2mks)**

……

……

……

……

b) Draw a similar diagram to show the effect of temperature on solid carbon (IV) oxide (dry ice) **(2mks)**

4) Identify the elements present in the following compounds.

i) Magnesium oxide. **(1mk)**

ii) Zinc chloride**.** **(1mk)**

iii) Sodium sulphate. **(1mk)**

iv) Aluminium nitrate. **(1mk)**

5) State three differences between luminous and non luminous flame. **(3mks)**

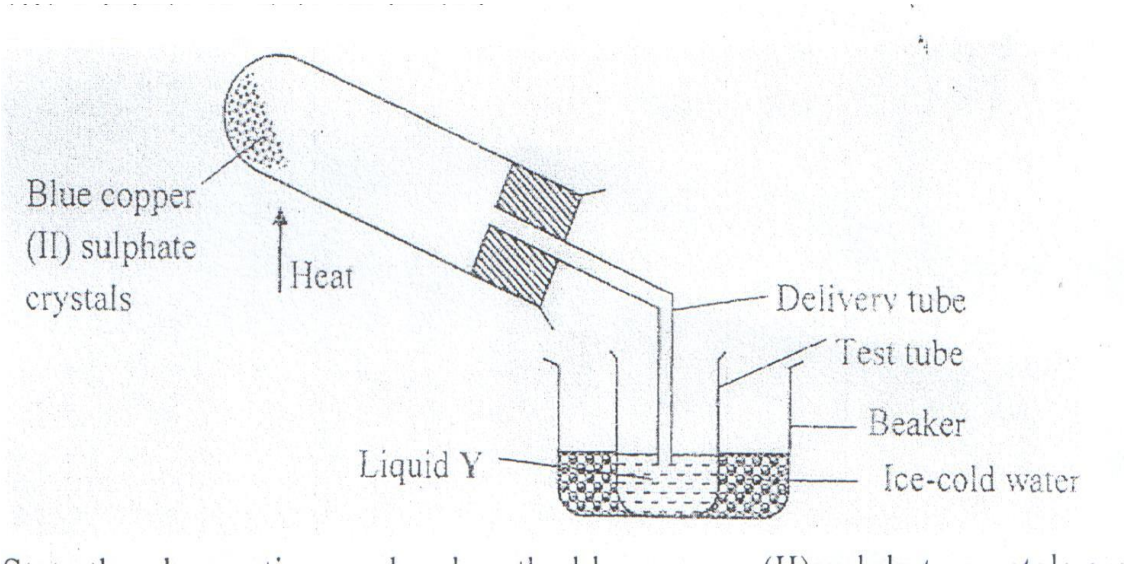
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6) The diagram below shows the effect of heat on hydrated copper (II) sulphate crystals.



a) State the colour of hydrated copper (II) sulphate crystals. **(1mk)**

……

b) What observations are made inside the boiling tube after heating has taken place? **(1mk)**

……

……

c) What is the use of ice-cold water in the beaker? **(½ mk)**

……

……

d) Name liquid Y. **(½ mk)**

……

e) What will be observed if the residue in the boiling tube is cooled and a few drops of liquid Y added to it? **(1mk)**

……

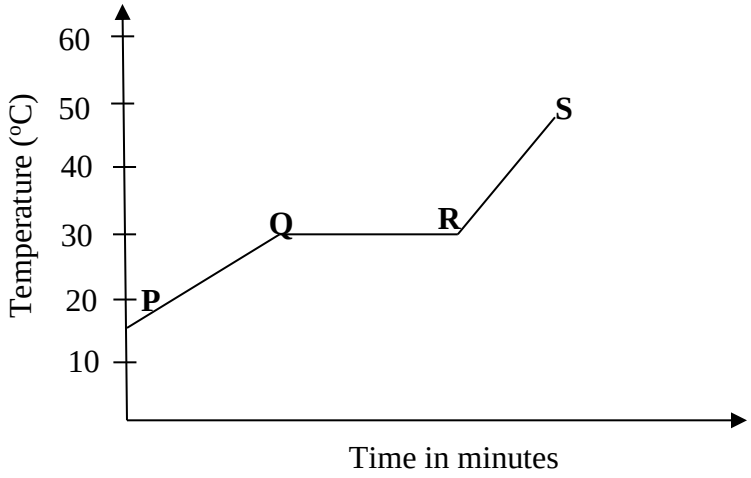
……

f) What type of chemical change is exhibited by copper (II) sulphate? **(1mk)**

……

……

7) The graph below shows the shape of the curve obtained by a student when solid X was heated to boiling.



a) Determine the melting point of solid X. **(1mk)**

……

……

……

b) After the experiment the student concluded that substance x was pure substance. Explain why he concluded so. **(1mk)**

……

……

c) If candle wax was used in the above experiment the portion QR would be horizontal. What does this tell us about candle wax? **(1mk)**

……

……

8) (a) What is the effect of impurities on the melting and boiling point of substances? **(2mks)**

……

……

……

(b) Explain the following:

i) During extraction of metals from their molten compounds for example calcium chloride is added to rock salt during the extraction of sodium from sodium chloride. **(1mk)**

……

……

ii) In temperate counties there is spreading of common salt in the roads during winter. **(1mk)**

……

……

9) Distinguish between temporary physical and permanent chemical changes. **(2mks)**

……

……

……

10) (a) Write the names of the elements represented by symbols. (3mks)

i) K –

ii) Cl –

iii) Fe –

iv) Mg –

v) C –

vi) Be –

(b) Define:

i) Atom. **(1mk)**

ii) Compound. **(1mk)**

iii) Element. **(1mk)**

11) (a) Complete the following. **(3mks)**

i) Acid + metal →

ii) Acid + Base →

iii) Acid + Carbonate →

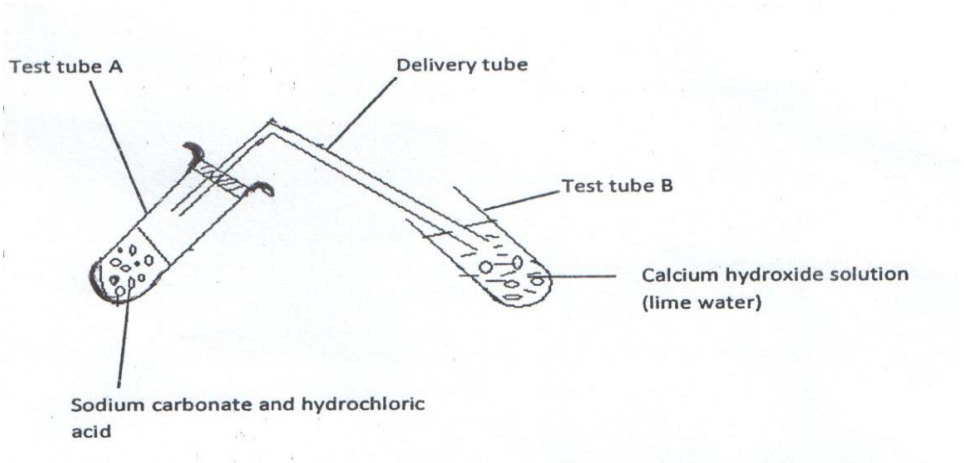
(b) Write word equations for the following reactions.

i) Zinc and hydrochloric acid. **(1mk)**

ii) Potassium hydroxide and sulphuric (VI) acid. **(1mk)**

iii) Magnesium carbonate and nitric (V) acid. **(1mk)**

12) Study the diagram below for the reaction between sodium carbonate and hydrochloric acid and answer the questions that follow.



(ii) Explain what was observed in:

i) Test tube A. (1mk)

ii) Test tube B. (1mk)

(iii) Write a word equation for the reaction which took place in test tube.

(1mk)

13) State whether solutions with the following pH values are acidic, basic or neutral. **(2mks)**

pH 3 –

pH 11 –

pH 14 –

pH 7 –

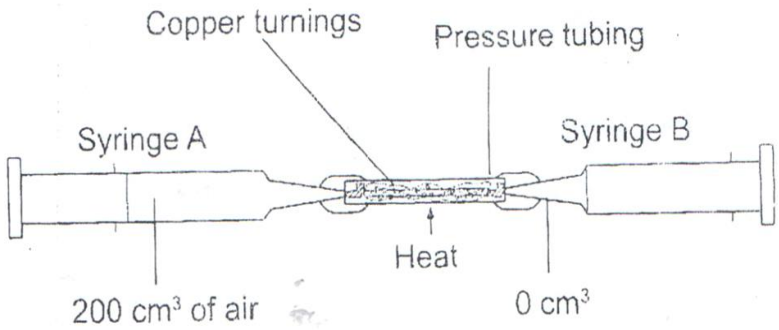
14) Is air a mixture or a compound? Explain. **(2mks)**

……

……

……

15) The apparatus below were used to determine the volume of oxygen in air. About 200cm^3 of air were passed repeatedly from syringe A to syringe B over heated copper turnings as shown in the diagram.



After sometime, the volume of air in the syringe A was 160cm^3 and syringe B, 0cm^3

a) Calculate the percentage of oxygen in the initial sample of ai**r.** **(2mks)**

……

……

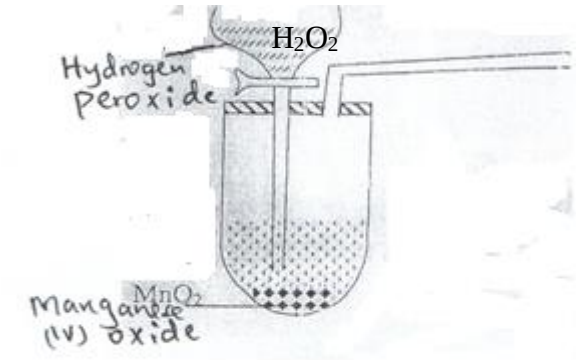
……

b) The percentage of oxygen calculated above was slightly less than the percentage of oxygen in air. Explain why. **(1mk)**

……

……

16) The set-up below was used to prepare a sample of oxygen gas. Study it and answer the questions that follow.



a) Complete the diagram to show how oxygen can be collected. **(2mks)**

b) Write a word equation for the reaction that produces oxygen above. **(1mk)**

……

……

c) State two physical properties of oxygen gas. **(2mks)**

……

……

d) State two uses of oxygen. **(2mks)**

……

……

17) Candle wax is an example of a hydrocarbon.

a) What is a hydrocarbon? **(1mk)**

……

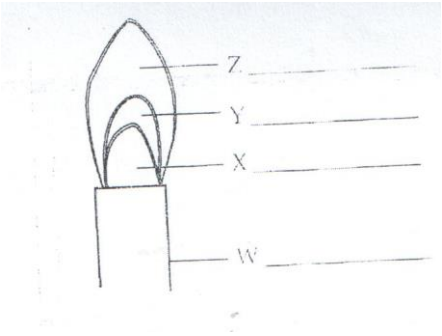
……

b) Name the two products formed when hydrocarbons burn in oxygen. **(1mk)**

……

……

18) Study the figure below and use it to answer the following questions.

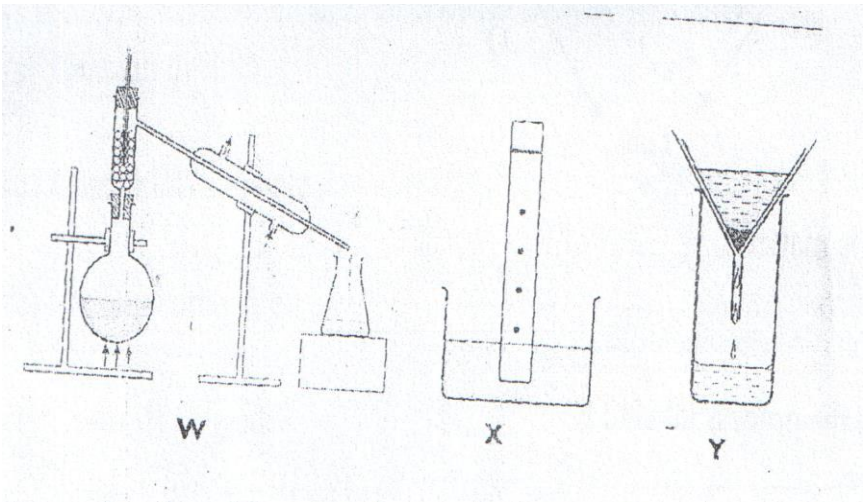


a) Name the parts labeled X and Y. **(1mk)**

b) Identify the flame. **(1mk)**

c) Which part of the flame is the hottest? **(1mk)**

19) The diagrams below show some physical methods of separating mixtures.



a) Name the methods of separation labeled W, X and Y.

……

……

……

……

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 6

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

QUESTIONS

1. (i). Define Chemistry**.** **(1mk)**

……

……

(ii) State three roles of chemistry in the society. **(3mks)**

……

……

……

2. (i) What is a drug. **(1mk)**

……

……

(ii)(a). Name three frequently abused drugs**.** **(3mks)**

……

……

……

(b). State two long term effects of drug abuse. **(2mks)**

……

……

3. (a). State three apparatus used for measuring accurate volumes of liquids. **(3mks)**

……

……

……

(b). Give two reasons why most laboratory apparatus are made of glass. **(2mks)**

……

……

4. What is a flame? **(1mk)**

……

……

5. Putting off flames is one of the laboratory safety rules. State three other rules**.** **(3mks)**

……

……

……

6. (i) What is a mixture? **(1mk)**

……

……

(ii). State two physical means of separating a mixture. **(2mks)**

……

……

7.Otieno, a form one student at Anestar High School accidentally mixed sulphur and iron filings.

(a). Suggest an appropriate method of separation you would advice him to use to separate the mixture. **(1mk)**

……

……

(b). Give a reason for the choice of your answer. **(2mk)**

……

……

(c). Describe how he would use the method named in (a) above to separate the sulphur and iron filings. **(3mks)**

……

……

8. (a) Name two substances that sublime when heated. **(2mks)**

……

……

(b). Give two reasons why dry ice (solid carbon (IV) oxide) is preferred to be used in cold boxes by ice cream vendors over ordinary ice. **(2mks)**

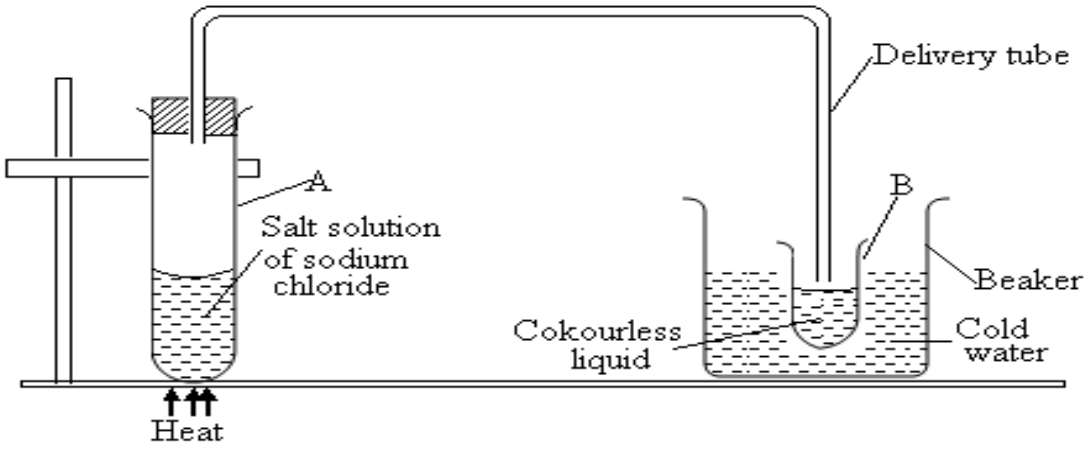
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……

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9. The set up below shows the apparatus used to demonstrate simple distillation process.

Study it and answer the questions that follow.



(a) What is observed in the delivery tube as the solution boils? **(1 mk)**

……

……

(b) What is the purpose of cold water in the beaker? **(1 mk)**

……

……

(c) What is observed in the test tube A after all the solvent has evaporated? **(1 mk)**

……

……

(d) Name the liquid you expect to collect in tube B. **(1 mk)**

……

……

10. Differentiate between physical and chemical changes as follows: **(4 mks)**

Chemical change	Physical change
(i)	
(ii)	
(iii)	
(iv)	

11. Write the chemical symbols of the following elements. **(5 mks)**

(i) Copper –

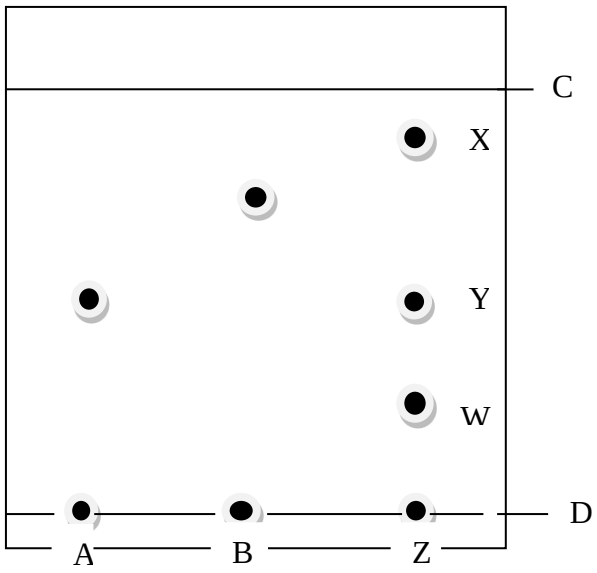
(ii) Sodium –

(iii) Potassium –

(iv) Lead–

(v) Calcium –

12. Sports of pure pigments A and B and a mixture Z were placed on a filter paper and allowed to dry. The paper was then dipped in a solvent. The results obtained were as on the paper chromatogram.



(a) Which line is the: **(2mks)**

(i) base line –

……

……

(ii) Solvent front –

……

……

(b) Which of the pure pigments was a component of Z? Explain. **(2 mks)**

……

……

(c) (i) Name a solvent that is used in paper chromatography. **(1 mk)**

……

(ii) Why is water not a suitable solvent in paper chromatography? **(1 mk)**

……

……

(d) Write a word equation for the reactions between:

(i) Carbon and oxygen - **(2 mks)**

……

……

12. (a). Define the following terms:

A saturated solution. **(1mk)**

……

……

Crystallization. **(1mk)**

……

……

(b). Give two industrial applications of crystallization as a method of separating soluble substances from their solutions. **(2mks)**

……

……

13 define the following terms. **(3mks)**

Atoms:.

……

……

Molecules.

……

……

Compound.

……

……

14. Name the elements present in the following compounds **(3mks)**

Sodium bromide:

Lead sulphate:

Potassium iodide:

15. State three observation made when a piece of sodium is placed on the surface of water**.(3mks)**

……

……

……

16. Give three application of chromatography. **(3mks)**

……

……

……

……

17. Give one advantage of commercial indicators over flower extract indicates. **(1mk)**

……

……

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 7

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

Answer All the questions in the spaces.

1) a) What is a drug**?** **(2 mks)**

……

……

b) What is drug abuse**?** **(2mks)**

……

……

c) One of the effects of drug abuse is hallucination. What does this term mean. **(2mks)**

……

……

d) Name three frequently abused drugs**?** **(3mks)**

……

……

……

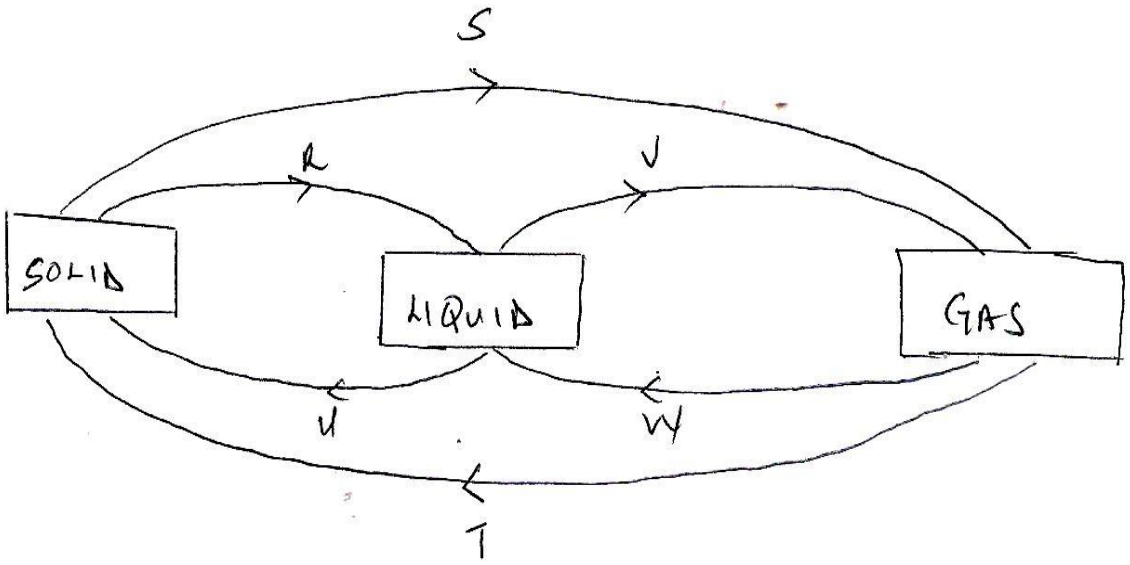
2) Distinguish between a conductor and a non-conductor and give and example in each**. (3mks)**

……

……

……

3) The diagram below shows the relationship between the physical states of matter.Study it and answer the questions that follows.



b) Identify the process R,V,W and **U** **(4 mks)**

……

……

……

……

c) Name three substances which can undergo the process represented by process S and **T.(3mks)**

……

……

……

4) The table below shows liquids that are miscible and those that are immiscible

Liquid	L3	L4
L1	Miscible	Miscible
L2	Miscible	immiscible

Use the information given to answer the questions that follow.

Name the method that can be used to separate L1 and L3 from a mixture of the two. **(1mk)**

……

Draw and name an apparatus that can be used to separate a mixture of L2 and L4**.** **(3mks)**

5) Give two reasons why most Laboratory apparatus are made of glass. **(2mks)**

……

……

6) Name three sources of heat beside Bunsen burner in the laboratory**.** **(3mks)**

……

……

……

7) a) Draw a labeled diagram of a non-luminous flame produced by the Bunsen burner**.** **(4mks)**

(iv) State two reasons why a non-luminous flame is preferred for heating**.** **(2mks)**

……

……

(v) After use a non-luminous flame should be put off or adjusted to a luminous flame. Explain.
(2mks)

……

……

8) Name three apparatus that are used to measure accurate volume of liquids. **(3mks)**

……

……

……

9) Distinguish between an element and a compound and give an example of each. **(3mks)**

……

……

……

10) By use of a diagram between a residue and a filtrate. **(2mks)**

11) Name the method you would use to separate the following mixtures.

Sand and ammonium chloride.(1mk)

……

Oil and Water. (1mk)

……

Kerosene and crude oil (1mk)

……

Salt and water.(1mk)

……

12) Describe how you would separate a mixture of salt,sand and iodine into different components.(**3mks)**

……

……

……

13) State the functions of the following apparatus as used in the laboratory.

Spatula **(1mk)**

……

……

Pine-clay triangle **(1mk)**

……

……

Wire gauze **(1mk)**

……

……

b Draw and state the use of a deflagrating spoon**.** **(3mks)**

……

……

……

……

14) State the two causes of accidents in a Chemistry laboratory**.** **(2mks)**

……

……

15) Define the following terms

Solvent extraction (2mks)

……

……

……

Hydrated salt (2mks)

……

……

……

Saturated Solution **2mks)**

……

……

……

16) State two functions of a fume cupboard as found in a chemistry **laboratory.** **(2mks)**

……

……

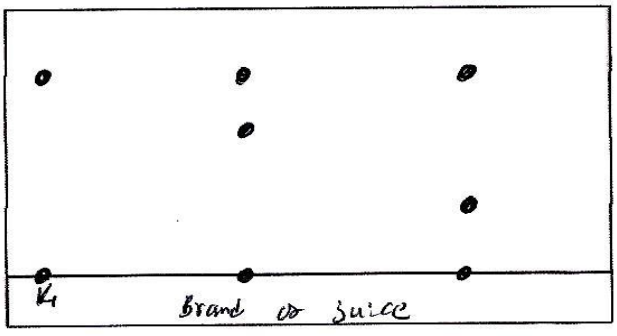
17) Explain the differences between solid and gaseous states using the theoretical model of matter in terms of the Kinetic **theory.** **(3mks**)

……

……

……

18) The diagram below represents a paper chromatogram for three brands of juices suspected to contain banned food colorings.



The results showed the presence or banned food colourings in L and M only.

On the same diagram

Circle the spots which show the banned food colourings. **(2mks)**

Show the solvent front. **(2mks)**

State two factors that determine the position where the pigments are deposited in the paper chromatogram from the point of origin. **(2 mks)**

……

……

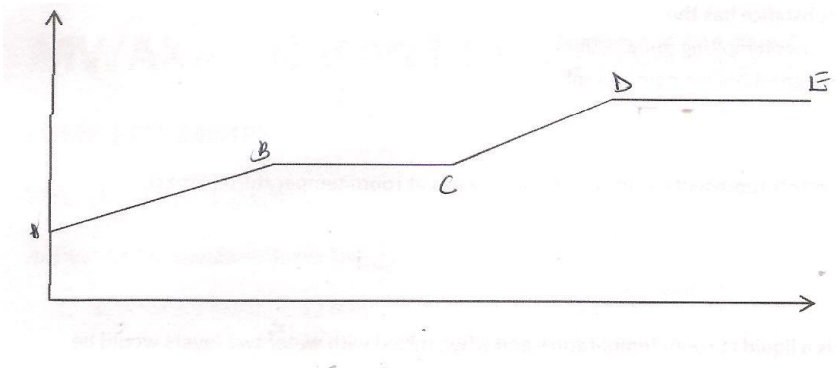
19) Classify the following processes as either chemical or Physical process type of change **(3mks)**

Heating copper(ii) sulphate crystals

obtaining Kerosene from crude oil

Souring of milk.

20) The figure below shows a heating curve of a certain pure solid.



What is happening at the stages represented by BC and CD (4mks)

……

……

……

……

On the diagram draw a heating curve of an improve substanc**e. (2 mks)**

21) Common table salt is contaminated with copper (ii)oxide. Explain how Pure sodium Chloride can be obtained from the mixture**. (3mks)**

22) The table bellows gives information on some substances. Use it to answer the question that follows.

Substances	Melting Point ^oC	Boliling point ^oC	Solubility in water
A	- 177	78.5	Very Soluble
B	-23	77	Insoluble
C	-219	-183	slightly soluble
D	-78	-33	Soluble

Which substance has the

Lowest melting **point** **(1mk)**

Highest boiling point **(1 mk)**

Which letters represents a substance that is a gas at room temperature**.** **(2mks)**

Which is a liquid at room temperature and when mixed with water two layers would be formed.(1 mk)

Which substance dissolves in water and could be separated from the solution by fractional distillation.(2mks)

23) a) Give the symbols of the following elements (3mks)

Sodium

Calcium

Potassium

20) Name the elements presents in the following compounds **(2mks)**

Zinc sulphide

Sodium oxide.

FORM 1 EXAMINATION TRIALS

CHEMISTRY

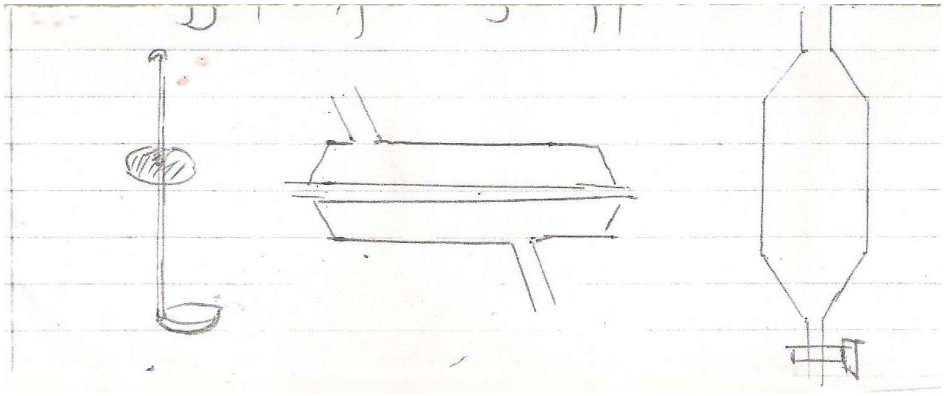
PRACTICE EXAM 8

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

1) Identify the following apparatus and their uses.(6mks)



A B C

Apparatus **uses** **6MKS**

A-

B-

C

2) State four preventive measure of drug –abuse. **(4mks)**

……

……

……

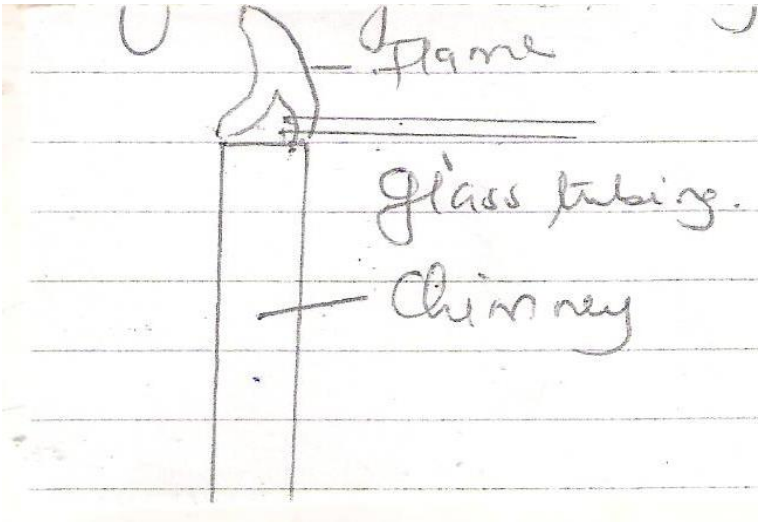
……

3) a.State under which conditions are the following flame formed in the laboratory. **2mks**

Luminous……..

Non- luminous……..

b.In an experiment, a form one student at Kiranja Secondary School placed an end of narrow glass tubing in the inner core of non-luminous flame and lit at the tip of the glass tubing as shown below .



State and explain the observation made at the tip of the glass tubing. (2mks)

4) State one application for each of the following methods of separating mixtures.

Filtration **(1mk)**

Fractional distillation**(1mk)**

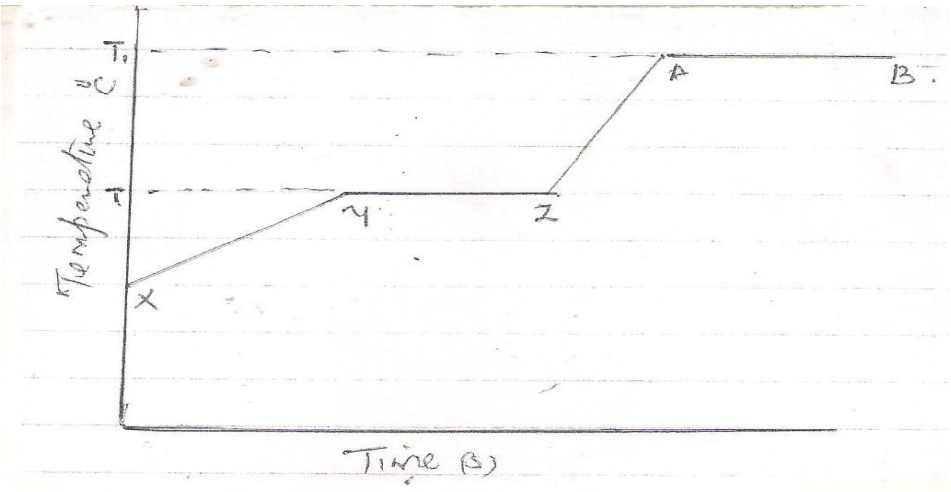
Solvent extraction**.(1mk)**

5) a.State two ways for determining the purity of substance**.** **(2mks)**

……

……

b.The diagram below represents heat curves of a pure surface of solid .Study it and answer the questions that follows.



Region YZ and AB have one thing in common .State it and explai**n.** **(2mks)**

……

……

ii.State the physical state of substance at the following region

XY. (2mks)

ZA (1mk)

6) Define the following terms.

Solution **1mks**

……

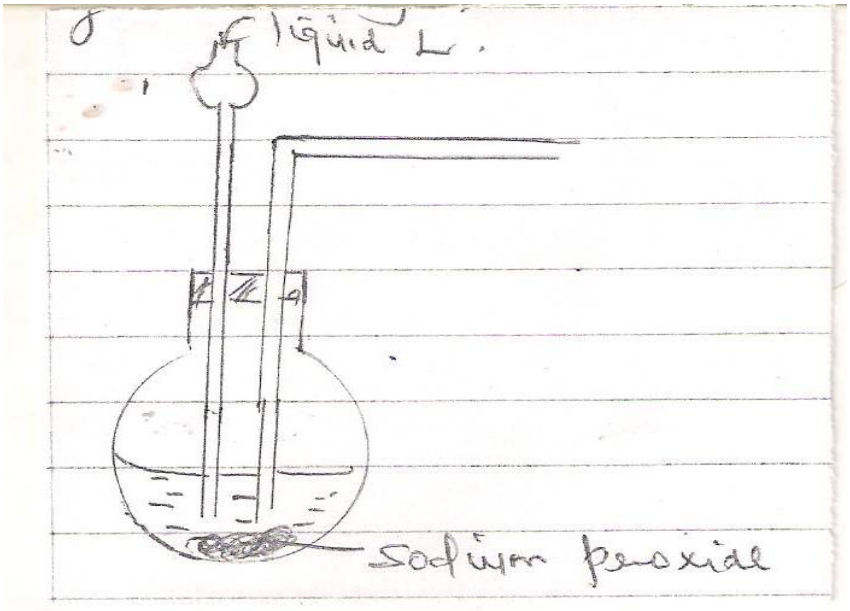
……

Saturated solution **1mks**

……

……

7) The following is a set up to prepare oxygen gas .Study it and answer the questions that follows.



Identify the mistake made when setting up the apparatus. **(2mks)**

……

……

If the mistake was corrected, complete the diagram showing how to collect dry oxygen gas. **(3mks)**

……

……

……

Identify liquid L . **(1mk)**

……

Write word equation for reaction at round- bottomed flask**.** **(2mks**

……

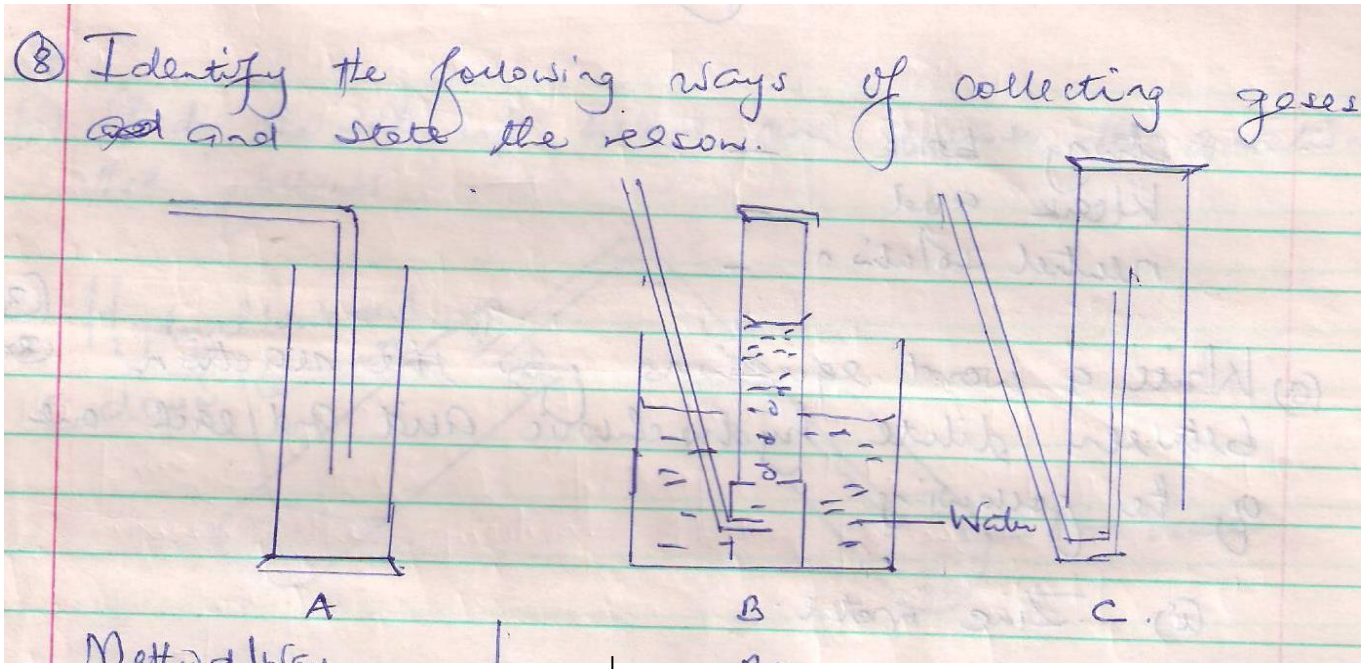
……

State two physical properties of oxygen**.** **2mks)**

……

……

Identify the following ways of collecting gases and state the reason.



Method/Way	Reason
A	
B	
C	

(3mks)

8) a. State one advantage of the universal indicator over acid- base indicator.(1mk)

b.The pH of the solutions are given as below

Solution	pH
A	7
B	6
C	2
D	12
E	8

Which of the pH listed above is **(5mks)**

i) Strong acid -

ii) Weak base –

iii) Strong base-

iv) Weak acid-

v) Neutral solution-

c.Write a word equations for the reaction between dilute hydrochloric acid and each one of the following

i. Zinc metal (**2mks)**

ii. Calcium carbonate **(2mks)**

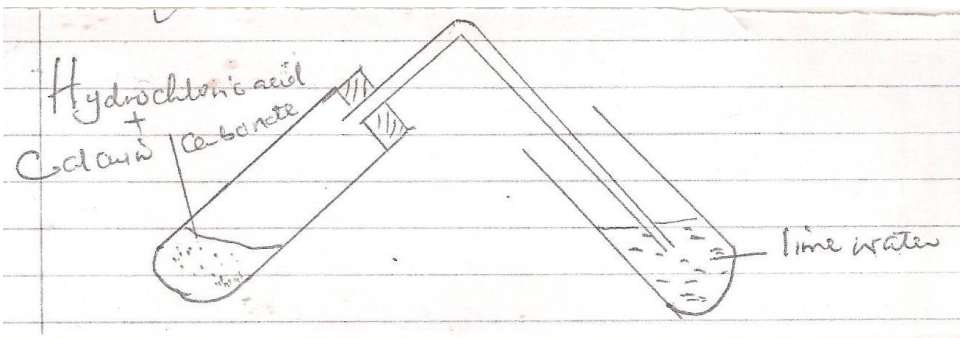
iii.Magnesium Oxide**(2mks)**

IV.Sodium hydroxide **(2mks)**

d.Miriam a form two student at Mutuma Trinity School was stung by wasp.Kanyotu advised her to apply sodium hydrogen carbonate .Explain. **(2mks**

e.Why is not advisable not to use sodium hydroxide**.(1mk)**

f.Study the set-up below and answer the questions that follows.



24) What are the products of the reaction between calcium carbonate and hydrochloric acid(give a word equation for the reaction**).** **(2mks)**

25) What would expect to observe in the limewater? **(1mk)**

g.i)State two uses of bases**.** **(2mk)**

……

……

ii)Two uses of acids. **(2mks)**

……

……

9) A student found a colourless liquid at laboratory. Describe two reagents he/she can use to determine the colourless liquid is wa**ter.** **(2mks)**

10) a.State two differences between permanent change and temporary physical changes.

Permanent change	**Temporary physical changes**

(4mks)

b.Study the following chemical equations (2mks)

d) Zinc oxide $\xrightleftharpoons{\text{heat}}$ zinc oxide

(white) yellow

e) Potassium $\longrightarrow$ Potassium + Manganeese (iv) + oxygen

Manganate(vii) Manganete(vi) oxide

f)Hydrated copper(ii) Oxide → ~~Copper~~ (ii) oxide + Water

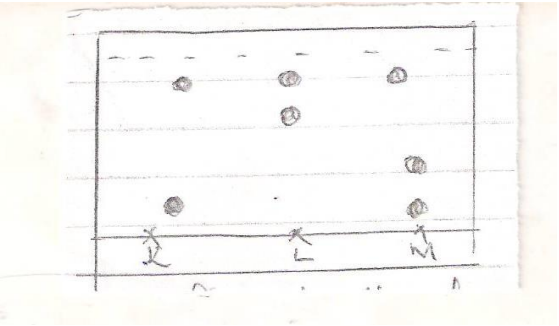
Identify the changes in (3mks)

Reaction I

Reaction II

Reaction III

11) The diagram below represents a paper chromatogram for three brands of juices suspected to contain banned food colourings.



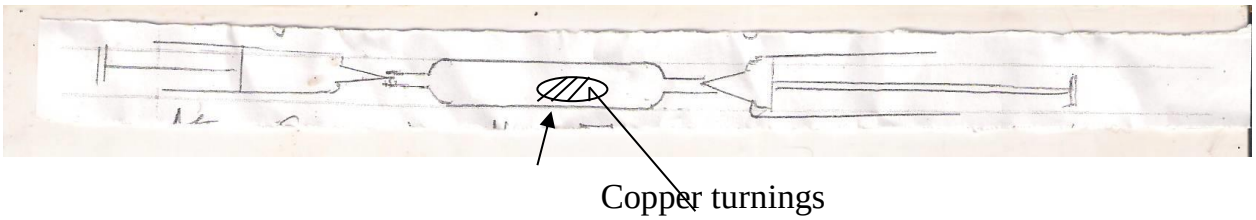
The results showed the presence of banned food colourings in L and M.One the same diagram.

v) Give the spots which show the banned food colourings.(2mks)

vi) Show solvent front.(1mk)

vii) State two applications of chromatography.(2mks)

12) The apparatus below were used to determine the volume of oxygen in air. About 200cm^3 of air were passed repeatedly and slowly from syringe A to syringe B, over heated copper turnings as shown in the diagram.



After sometime the volume of air syringe A was 160cm^3 and syringe B 0cm^3.

(i) Calculate the percentage of oxygen in the initial sample of **air.** **(2mks)**

(ii) Write down a word equation for the reaction that took place in the combustion **tube.** **(1mk)**

(iii) What are possible sources of error in the experiment**.** **(2mks)**

(iv) Explain why the air is passed slowly and repeatedl**y.** **(1mk)**

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 9

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

QUESTIONS

1. a) Define the following terms**.** **3mk**

I)Indicator

II) Neutralization

III) Alkali

b) The table below shows the pH values of some solutions. Study it and answer the questions that follow.

solution	PH
M	9
R	7
K	3
D	13
G	6

Which solution is likely;

i) To produce bubbles/ effervescence with sodium carbonate solution. **1mk**

……

Ii) To be sodium chloride solutions **1mk**

……

c) Choose two solutions that can react to form a salt and water only**.** **1mk**

……

……

d) Which color would solution M show when added methyl orange indicat**or.** **(1mk)**

……

……

e. Give a reason why the flower extract indicator is not **preferred.** **1mk**

……

……

f/. What is the advantage of the universal indicator over the other indicators**.** **1mk**

……

……

2. a) Indicate whether each of the following is a mixture or a compound**.** **(6mk)**

I) Black ink----------

Ii) Sodium chloride------

Iii) Tea----------

Iv) Crude oil----------

v) sea water----------

vi) Air----------

b) State any four differences between a permanent and a temporary change. (**4mk.)**

Permanent	Temporary

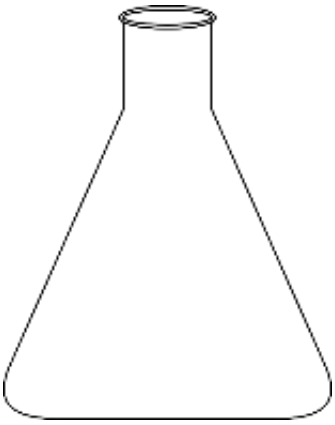
C) A blue compound was heated in the lab in a boiling tube, droplets of a colorless liquid were seen on the cooler parts of the tube and the solid turned to white.

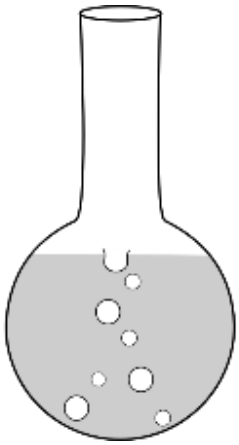
I) Name the type of change undergone by the substance **above** **(1mk)**

II) Identify; _ the white solid **(2mks)**

_the colorless liquid

3.(a)Name the following apparatus (**2mks)**

i) 

ii) 

b) Draw the shapes of the following lab apparatus; **(4mks)**

i) Deflagrating spoon ii) Pipette

iii) Teat pipette iv) Filter funnel

c) State the function of the following lab apparatu**s;** **(2mks)**

i) Dropping funnel

Ii) Desiccator

d) List two apparatus used to measure accurate but specific volumes of liquid. **(2mks)**

4. a) What is a flame? (1MK)

b) List any four differences between a luminous and a non-luminous flame**.** **(4mks)**

……

……

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……

C) Explain why the luminous flame of the Bunsen burner produces a lot of light **(1mk)**

D) State the functions of the following parts of the Bunsen burner **(2mks)**

i)The collar---

ii) The base ------

e) Which type of flame is most preferred for heating and why? Give two reasons it's preferred. **3mk)**

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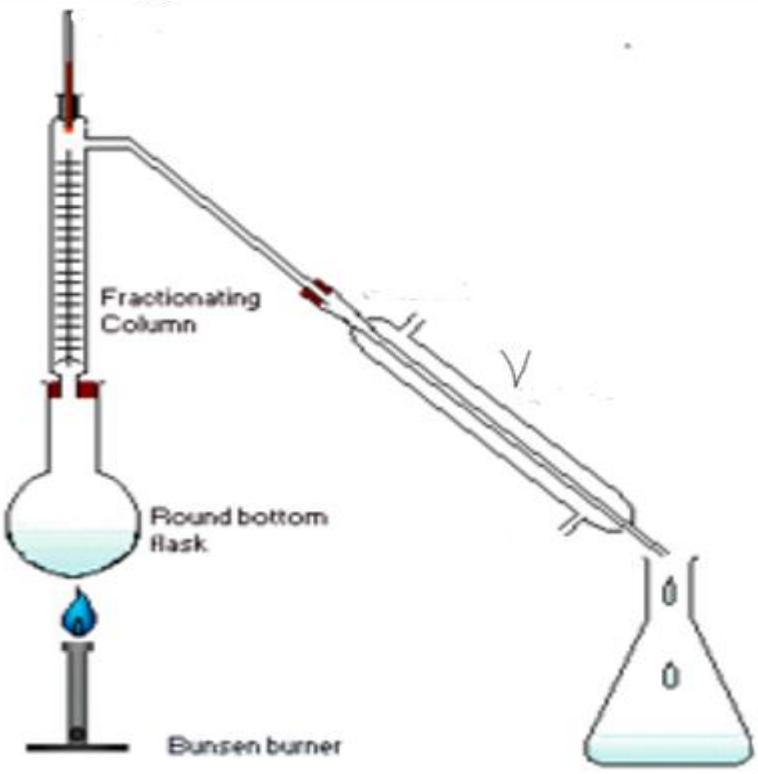
5. a) Choose the most suitable **method** to get the ***first*** substance from the following mixtures;
(3MK)

i) Water and salt----------

ii) Sand and water ----------

iii) Oil and milk----------

b) A student was given a mixture of solution x and solution y which were miscible with boiling points of 84^0c and 75^0c. The student's setup the set of apparatus as shown below to separate the two liquids. Study it carefully and answer the questions that follow



i) Name the method above that the student used **(1mk)**

ii) What name is given to apparatus V **(1mk)**

iii) Indicate on the diagram the direction of flow of **water** **(1mk)**

iv) State the function of each of the following**;** **(2mks)**

I)Fractionating column-

II) Thermometer –

v) Which of the two liquids was collected as the first distillate and why**?** **(2mk)**

6. Draw a well labelled diagram to show how a mixture of sodium chloride and iodine can be separated to acquire a pure sample of each**.** **(3MK)**

7. **DESCRIBE** how an oil sample can be extracted from ground nuts. **(3mks)**

8 (a). FILL in the table below. **(8MK)**

ELEMENT	SYMBOL
Calcium	
	Cu
	K
Sodium	
	F
Helium	
Silver	
	Mg

b) Identify the elements present in each of the following compounds;

i) Lead chloride **(1mk)**

ii) Zinc carbonate **(1 ½ Mk)**

iii) Aluminum **sulphate** **(1 ½ Mk)**

9. STATE and EXPLAIN the changes in mass that occurs when each of the following substances are strongly heated in a boiling tube in the lab. **(2mk)**

a) Copper Nitrate

b) Zinc oxide

10. Define the following terms; **(6mks)**

a) Chemistry

……

……

b) Drug

……

……

c) Drug abuse

……

……

d) Molecule

……

……

e) Atom

……

……

f) Element

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……

11. Fill in the gaps in the chemical equations below.

a) Zinc + hydrochloric acid ⟶ ---------- + Hydrogen

(1mk)

b) Copper ii oxide + Sulphuric (vi) acid ⟶ ---------- + ----------

(2mk)

FORM 1 EXAMINATION TRIALS

CHEMISTRY

PRACTICE EXAM 10

TIME: 2 HOURS

NAME…….. ADM NO…….....

STREAM……..……. DATE;……...

Answer All the questions in the spaces.

1. What is chemistry **(1 mark)**

……

……

2. State any importance of studying chemistry **(1 mark)**

……

……

3. **a)** Name three illegal drugs **(3 marks)**

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……

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b) State three ways of preventing drug abuse. **(3 marks)**

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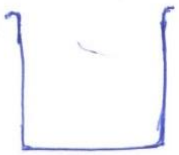
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4.

Complete the following table **(8mks)**

Apparatus	name	use
i 		
ii 		
iii 		
iv 		

5. State four laboratory safety **rules** **(4mks)**

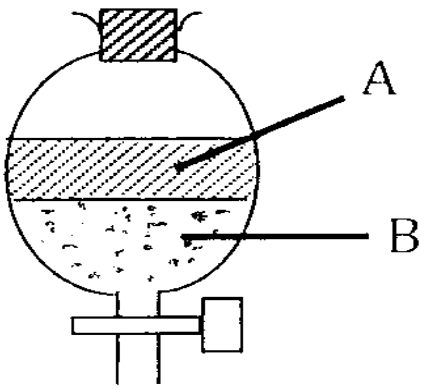
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6. (a)The apparatus below were used to separate a mixture of liquid A and B.



State ***two*** properties of liquids that make it possible to separate using such apparatus. **(2 marks)**

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Give the name of the above apparatus. **(1 mark)**

……

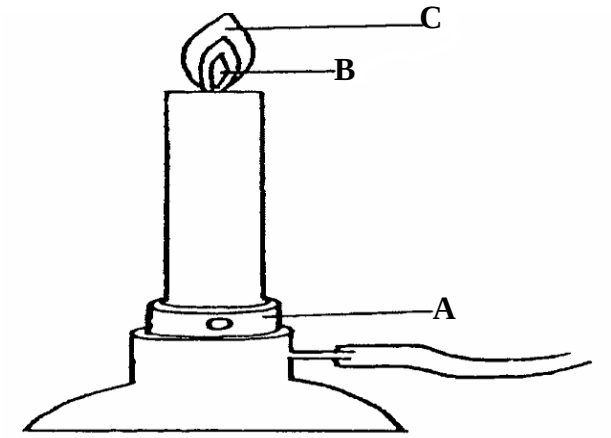
7. Describe how you can separate a mixture of sand and common salt **(3 marks)**

……

……

……

8. The diagram below shows a Bunsen burner when in use



Name the regions labelled **B** and **C**. **(2 marks)**

B…….

C…….....

What is the function of the part labeled **A**? **(1 mark)**

9**.** State three differences between physical and chemical change. **(3 marks)**

……

……

……

10. **i.** Differentiate between weak and strong alkali **(2 marks)**

……

……

ii. The following is a list of pH values of some substance:

Substance	**M**	**N**	**V**	**X**	**Z**
pH	10.6	7.2	13.2	5.9	1.5

Identify:

a) Strong acid **(1 mark)**

b) Weak base **(1 mark)**

11. Name the method or process that can be used to separate each of the following mixture

d) Water and motor oil **(1 mark)**

ii . Iron filings and sulphur powder . **(1 mark)**

(v) Iodine and sand **(1 mark)**

iv. Water and ethanol **(1 mark)**

12. Write the chemical symbols of the following elements **(5 marks)**

13) Calcium ……..

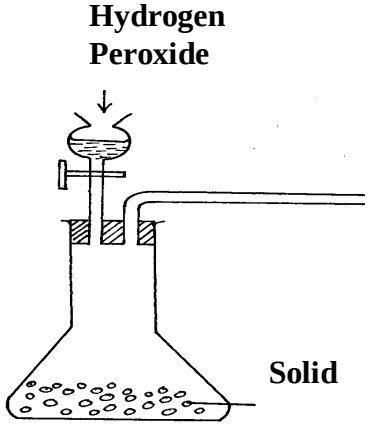
14) Sodium…….

15) Barium……..

16) Lead……

17) Copper…….

13. The diagram below represents the apparatus used to prepare and collect oxygen.



viii) Complete the diagram to show how oxygen gas is collected. **(2 marks)**

ix) Name solid **P** **(1 marks)**

P

iii. Write the word equation to show the reaction that produces oxygen gas. **(1 mark)**

x) State three physical properties of oxygen. **(3 marks)**

v. State **two** uses of oxygen gas. **(2 marks)**

14.(i) What is a flame? **(1mk)**

(ii). The following diagrams represent the two types of flames produced by a bunsen burner.

(a) (b)



(a). Identify the flames (a) and (b). (2mks)

(b). Which type of the flames identified above is preferred for heating? Give a reason for your answer.

(2mks)

(c). give four differences between the flames (a) and (b) above. **(4mks)**

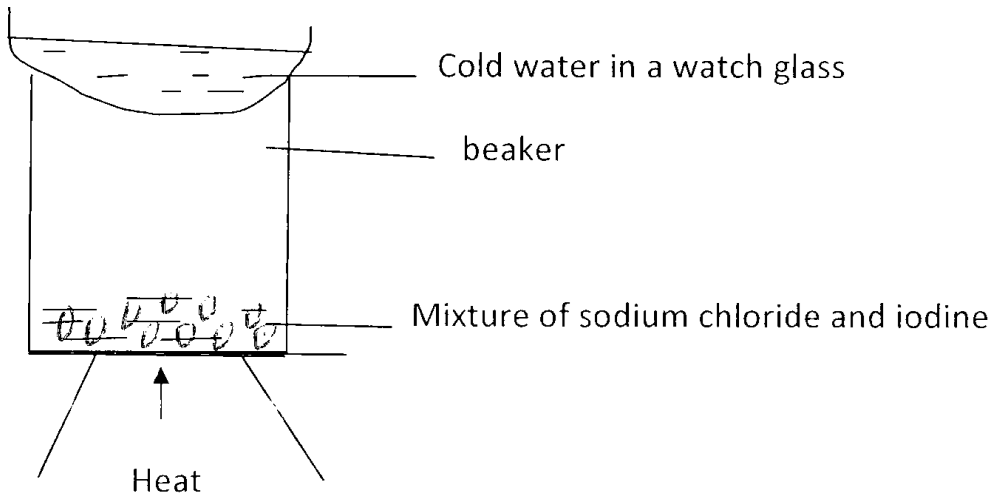
Flame (a)	Flame (b)

(d)(i). Give two reasons why flames (a) and (b) in 4(ii) above differ. **(2mks)**

……

……

15. A form one student wanted to separate and obtain iodine and sodium chloride (common table salt) from a mixture of the two. He set the experiment set up shown below.



(a). the mixture was heated for some time and left to cool. On cooling, shiny black crystals and white crystals were observed on the surface of the watch glass and in the beaker respectively. Name:

I. Shiny black crystals

II. White crystals. **(2mks)**

(b). What was the purpose of the cold water in the watch glass**?** **(1mks)**

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……

(c). What property of iodine makes it be collected on the watch glass as shown? (1mk)

……

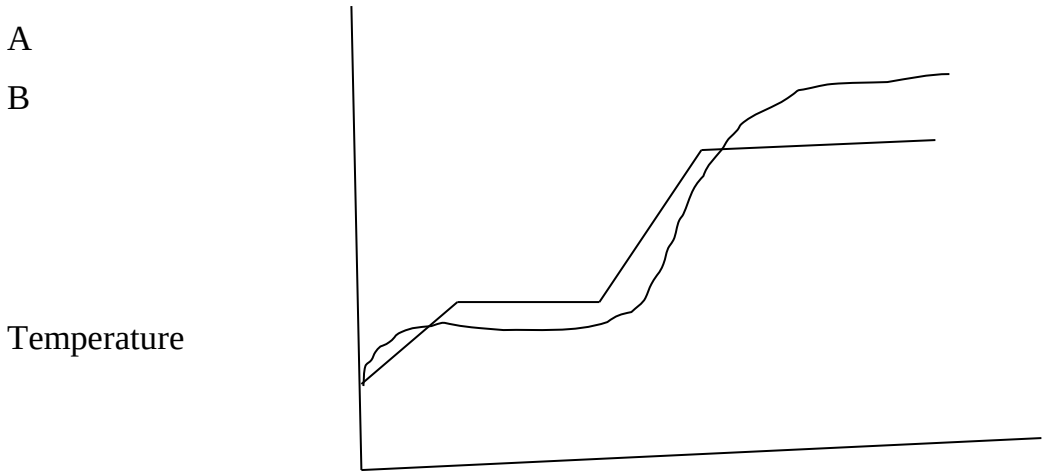
……

(d). Explain why it is possible to separate a mixture of iodine and sodium chloride.(1mk)

……

……

16. The curves below represent the variation of temperature with time when pure and impure samples of a solid were heated separately.



(a). Which curve represents the variation in temperature for pure solid? Explain. **(2mks)**

……

……

(b)State the effect of an impurity on the melting and boiling points of a pure substance**.(2mks)**

……

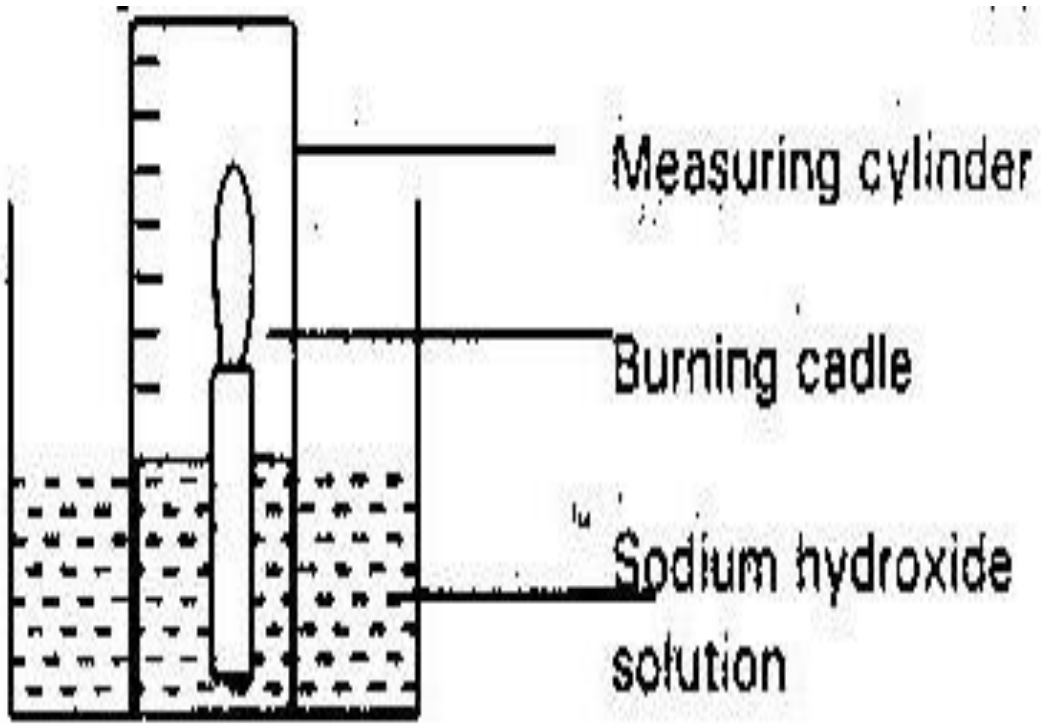
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(c) Name two gases used with oxygen in welding. **(2mks)**

……

……

17. A candle was burnt using the apparatus shown below. The initial volume of measuring cylinder was $90cm^3$. The apparatus was allowed to cool and the volume of air in the measuring cylinder had dropped to $70cm^3$.



a) Why was the volume recorded when the air was cooled? **(1mk)**

……

……

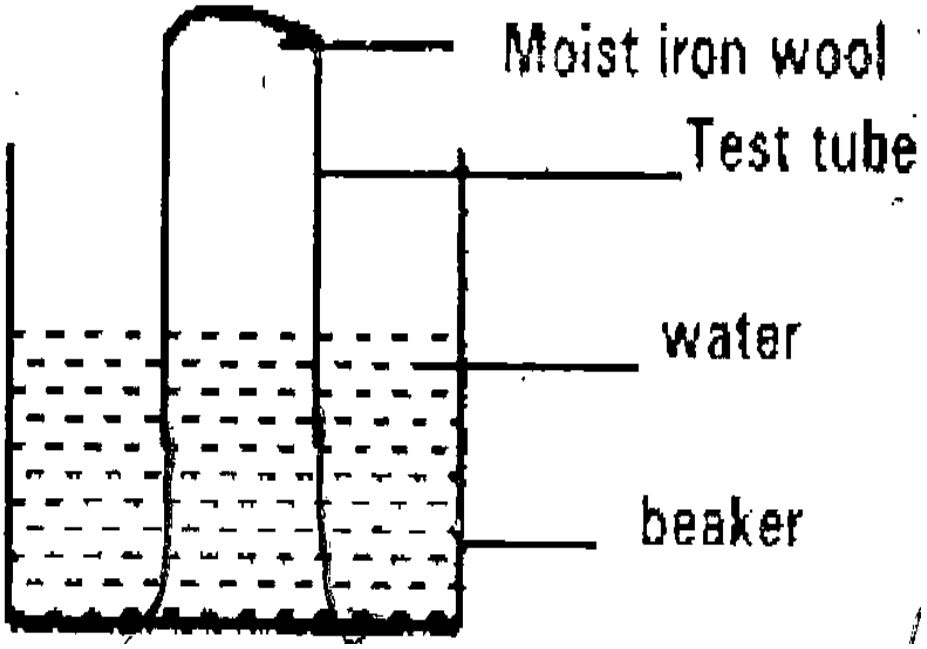
b) What was the purpose of sodium Hydroxide? **(1mk)**

……

……

c) Use the results given to calculate the percentage of oxygen in air. **(2mks)**

18. The set up below was used to study some properties of air



State and explain two observation that would be made a t the end of the experiment. **(3mks)**

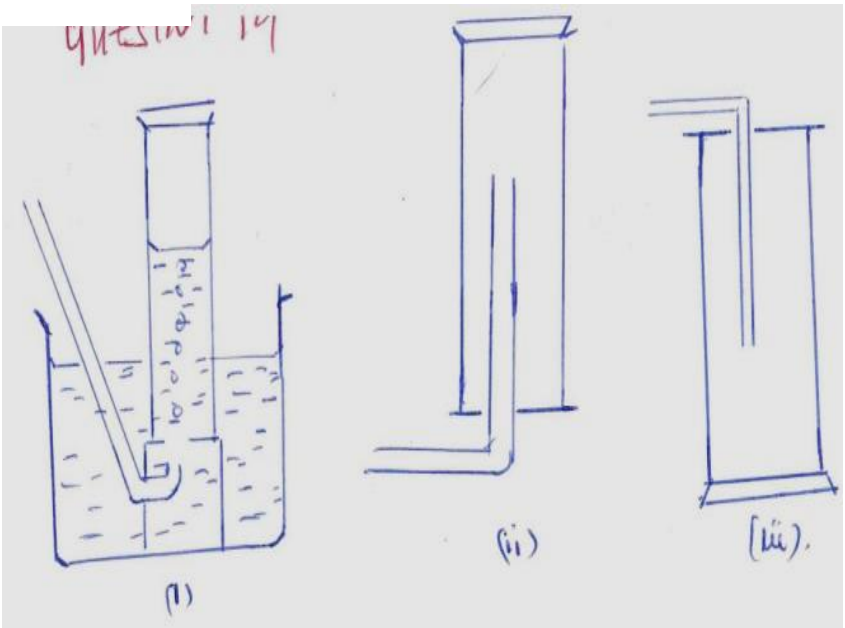
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19. The diagram below represents three methods for collecting gases in the laboratory



a) Name the methods shown in the diagram **(3mks)**

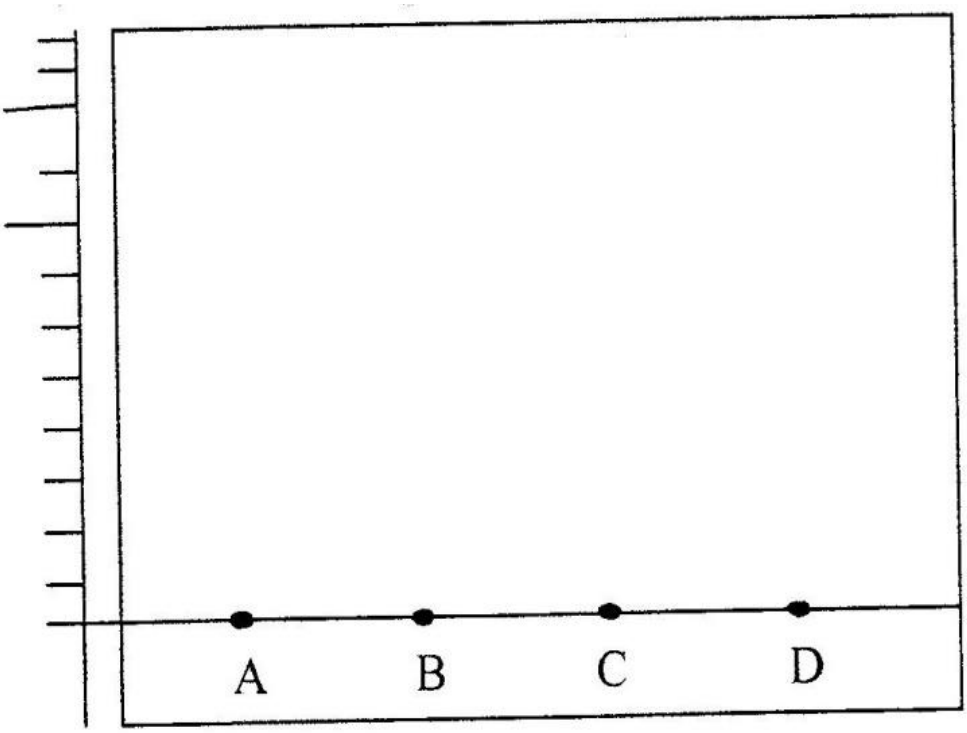
b) State with reasons the most suitable methods for collecting each of the following gases.

i) Oxygen (1mk)

ii) Hydrogen (1mk)

iii) Carbon (IV) Oxide (1mk)

20 a) The diagram below shows spots of pure substance A,B, and C on a chromatography paper. Spot D is that of a mixture



After development, A, B and C were found to have moved 8cm, 3cm and 6 cm respectively.

D has separated into two spots which had moved 6cm and 8 cm

(i) On the diagram

I Label the baseline (origin) **(1 mark)**

II Show the positions of all the spots after development **(3 marks)**

(ii) Identify the substances present in the mixture D **(2 marks)**

(b) Describe how solid ammonium chloride can be separated from a solid mixture of ammonium chloride and anhydrous calcium chloride **(2 marks)**

……

……

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……

(c) The table shows liquids that are miscible and those that are immiscible

Liquid	L_3	L_4
L_1	Miscible	Miscible
L_2	Miscible	Immiscible

Use the information given to answer the questions that follow

(i) Name the method that can be used to separate L1 and L3 from a mixture of **two** **(1 mark)**

……

……

(ii) Describe how a mixture of L_2 and L4 can be separated **(2 marks)**

……

……

……

FOR THE FOLLOWING;

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