Select the best answer.

Silicon is an element that is part of the p-block of the periodic table. Silicon has four valence electrons. Knowing this, which of these must be silicon's electron configuration?



ATTEMPT 1

- ◉ $[Ne]\ 3s^2\ 3p^2$ ✓
- ○ $[Ne]\ 3s^2 3p^4$
- ○ $[Ar]\ 4s^2 3d^4$
- ○ $[Ar]\ 4s^2 3d^2$

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Select the best answer from the drop-down menu to complete the sentence.

On the periodic table, the 1. s-block ✓ includes the table's most reactive metals.

1. s-block ✓

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Select the best answer from the drop-down menu to complete the sentence.

The f-block of the periodic table is comprised mainly of 1. radioactive metals ✓.

1. radioactive metals ✓

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Select the best answer.

The electron configuration for copper is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1\ 3d^{10}$ How does its configuration make it different from most other d-block elements?



- Copper has three completely filled *s* orbitals.
- Copper has a completely filled *d* orbital in its second highest energy level.

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- ✓ Copper only has one electron in its outermost *s* orbital.
- Copper only has one orbital that is not completely filled.

Select the best answer.

Which of these elements is found in the periodic table's d-block?

- [Ne] $3s^2\ 3p^6$
- [Kr] $5s^1\ 4d^5$ ✓
- [Xe] $6s^2\ 5d^1\ 4f^{14}$
- [Ar] $4s^2\ 3d^{10}\ 4p^2$

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Drag and drop the best answers to complete the sentence.

In the f-block of the periodic table, actinides have partially filled f orbitals in the

1. fifth ✓ principal energy level, while lanthanides have partially filled f orbitals in the 2. fourth ✓ principal energy level.

first | second | sixth | third

1. fifth ✓

2. fourth ✓

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