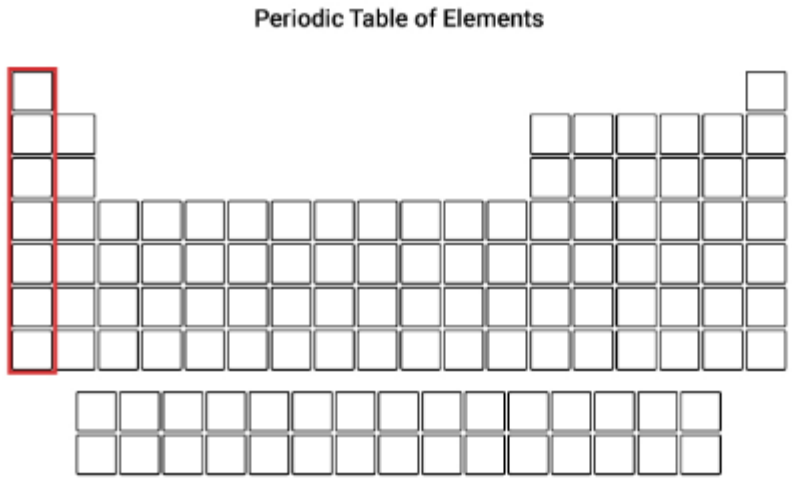
Fill in the blank with the correct term from the Word Bank.

Word Bank			
1+	2+	1-	2-



The elements located in this column of the periodic table

form monatomic ions with a

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1. 1+ ✓ charge.

Select the best answers from the drop-down menus to complete the sentences.

S: $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^4$

The electron configuration for sulfur (S) shows that there are [1. 6 ✓] electrons in the outer shell. Sulfur gains [2. 2 ✓] electrons when it becomes a sulfide (S^{2-}) ion and has [3. 8 ✓] electrons in its outer shell.

1. 6 ✓
2. 2 ✓
3. 8 ✓

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Select the best answer.

Which of the following **best** describes the relationship between the charge of a monatomic ion and its oxidation number?

ATTEMPT 1

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- The oxidation number is the number of electrons an atom can share.
- The oxidation number is the negative charge of a stable ionic compound.
- ✓ The oxidation number is the charge of a monatomic ion.
- The oxidation number is the energy level of the valence shell.

Select the best answer.

Which of the following statements best describes why transition elements have multiple oxidation numbers?

- ◉ Transition elements can lose electrons from the d orbitals. ✓
- ○ A transition element does not need a full valence shell of electrons to be stable.
- ○ A transition element can have fifteen of electrons in the innermost electron shell.
- ○ Transition elements have characteristics of metals and nonmetals.

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Select the best words from the drop-down menus to complete the sentence.

When sulfur gains two electrons, it forms 1. an anion ✓ with a 2. 2− ✓ charge.

1. an anion ✓

2. 2− ✓

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Select the best answer.

Which of the following explains what happens when an atom becomes a cation?

- ◉ The atom loses one or more electrons to form a positive charge. ✓
- ○ The atom gains one or more electrons to form a positive charge.

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- ○ The atom loses one or more electrons to form a negative charge.
- ○ The atom gains one or more electrons to form a negative charge.