9. Which of the following is correct about the reaction at Equilibrium?
   A) The rate of forward and backward reaction are equal
   B) The reaction quotient is greater than equilibrium Constant
   C) The System is static in nature
   D) The concentrations of the reactants and Products are Equal

10. A type of Solution which Contain less amount of solute than solvent is Known as_____.
   A) Super Saturated Solution C) Saturated Solution
   B) Unsaturated Solution D) Concentrated Solution

11. Molecules that have the same molecular formula but different Structural formulas are called _____.
   A) Stereoisomers B) Allotropes C) Isotopes D) Structural isomers

12. From the following chemical bond one is formed by sharing of electron between two elements.
   A) Ionic Bond B) Metallic Bond C) Covalent Bond D) Vander waal forces

13. Among the following acid, which one is found in car battery as Electrolyte?
   A) H_2SO_4 B) HCl C) HNO_3 D) H_3PO_4

14. A process of directly converting from solid phase to gaseous phase is known as_____
   A) Sublimation B) Melting C) Condensation D) Deposition

15. What is the mass of 30.0 mL of liquid that has density of 1.60 g/mL
   A) 18.8g B) 48.0 g C) 31.6 g D) 53.3 g

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ARBA MINCH UNIVERSITY
STREAM SELECTION QUALIFICATION EXAMINATION
Chemistry for Health, Exam Code 03

Instruction: Choose the best answer and put the letter of your choice on the answer sheet

1. Which of the following is not colligative properties of Solutions?
   A) Boiling point of Solution
   B) Freezing point Depression
   C) Solution Saturation
   D) Osmotic Pressure
2. Molecules that have the same molecular formula but different Structural formulas are called _____.
   A) Stereoisomers B) Allotropes C) Isotopes D) Structural isomers
3. Which method of water treatment is useful for controlling disease causing Organisms Such as viruses, bacteria and Parasites?
   A) Water Softening
   B) Filtration method
   C) Chlorination Method
   D) Distillation Method
4. Up on dissolving common salts in water, it is observe that _____
   A) Freezing point is raised
   B) Boiling point of Solution Decreases
   C) Boling point of Solution increases
   D) Vapor Pressure is Increased
5. If a sample of matter is uniform throughout and cannot be separated into other substance by physical means, it is _____
   A) A compound
   B) Homogenous Mixture
   C) an Element
   D) Heterogeneous mixture
6. How many grams of NaOH would be needed to make 0.750 mol/L of Solution? (Mwt of NaOH = 40 g/mol)
   A) 1.50 g B) 0.150 g C) 150 g D) 15 g
7. What mass of $CaCl_2$ is required to prepare 500.0 mL of a 0.200 M Solution of $CaCl_2$ (Mwt of Ca = 40 g/mol and Mwt of Cl = 35.5 g/mol)
   A) 0.200 g B) 0.100 g C) 22.2 g D) 11.1 g
8. Which of the following is usually found in the Nucleus?
   A) Proton and Neutron
   B) Proton, Neutron and Electron
   C) Neutron only
   D) Electron and Neutron

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