



EDUCATIONAL ASSESSMENT AND EXAMINATIONS SERVICE
ETHIOPIAN SECONDARY SCHOOL LEAVING CERTIFICATE
EXAMINATION (ESSLCE)

Number of Item: 80
SUBJECT CODE_ 05

CHEMISTRY HAMILE, 2015 E.C (JULAY 2023 GC)

TIME ALLOWED 2^{1/2} HOURS
BOOKLET CODE_ 496

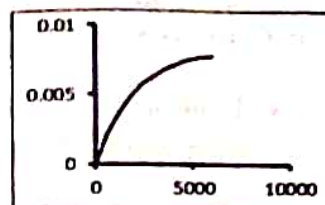
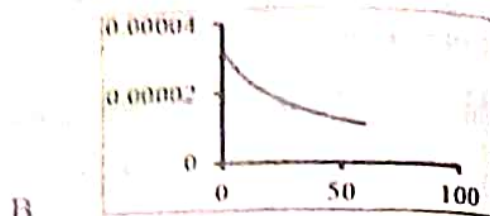
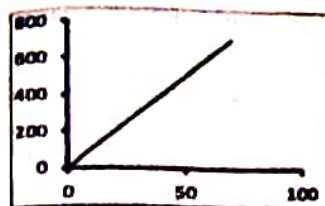
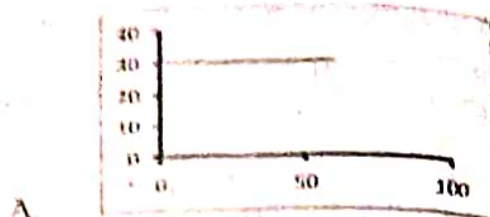
1. In a kinetic experiment, a student placed crystals of iodine in a closed reaction vessel, introduced a given quantity of hydrogen gas, and collected data to calculate the rate of formation of hydrogen iodide. What would be the effect obtained on the rate of the reaction, if the student takes more amount of iodine from that used in the first experiment?
- A. It speeds up the rate of the reaction.
 - B. It lowers the rate of the reaction.
 - ☒ C. The formation of hydrogen iodide will be decreased.
 - D. The rate of the reaction will be constant.
2. Which of the following **CORRECTLY** relates the vapor pressure of a liquid with its boiling point? Boiling point is
- A. the temperature at which all the three physical state co-exist.
 - B. the temperature at which the vapor pressure becomes equal to the external pressure.
 - ☒ C. the temperature at which molecules present in the vapor phase go back to the liquid state.
 - D. the temperature at which the kinetic energy of molecules is much less than the intermolecular force.
3. Consider this equilibrium system:
- $$\text{CO (g)} + \text{Fe}_3\text{O}_4(\text{s}) \rightleftharpoons \text{CO}_2(\text{g}) + 3\text{FeO}(\text{s})$$
- Which of the following disturbances shift the equilibrium position to the right?
- | | |
|-------------------------------------|--|
| A. Removing $\text{CO}_2(\text{g})$ | C. Removing $\text{FeO}(\text{s})$ |
| B. Removing $\text{CO}(\text{g})$ | ☒ D. Adding $\text{Fe}_3\text{O}_4(\text{s})$ |

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4. Which one of the following resulted from interaction of the atomic orbitals of bonding atoms to produce new electron distribution associated with the entire molecule?
- A. Hybrid orbital
B. Molecular orbital
C. Electron orbit
~~D. Bohr orbit~~
5. Which type of organic substances would be produced by heating a mixture of a carboxylic acid and an alcohol in the presence of H_2SO_4 ?
- A. Ketones
B. Aldehydes
~~C. Esters~~
D. Ethers
6. Which one of the following is the reducing agent in the following reaction?
- $$8\text{NH}_3(\text{g}) + 6\text{NO}_2(\text{g}) \rightarrow 7\text{N}_2(\text{g}) + 12\text{H}_2\text{O}(\text{l})$$
- ~~A. NO_2~~
B. NH_3
C. H_2O
~~D. N_2~~
7. Carbon forms two stable compounds with oxygen, namely, carbon monoxide and carbon dioxide. For a fixed mass of carbon, the ratio of oxygen in carbon monoxide to oxygen in carbon dioxide is 1:2. This result is consistent with the law of
- ~~A. definite proportion.~~
B. conservation of mass.
C. constant composition.
D. multiple proportion.
8. When a system at equilibrium is disturbed by a change of variable, the system shifts the equilibrium position in a way that tends to counteract this change of variable. This is a statement for
- ~~A. Pauli's exclusion principle~~
B. Aufbau principle
C. Heisenberg uncertainty principle
D. Le Chatelier Principle

9. Which one of the following graphs illustrates the rate change in terms of change in reactant concentration with time?



10. Which one of the following gas laws states the volume of a fixed amount of gas maintained at constant pressure is directly proportional to the absolute temperature of the gas?

A. Charles' law
B. Boyle's law

C. Avogadro's law
D. Combined gas law

11. The pH of a 0.1 M solution of formic acid (HCOOH) is 3. What is the K_a of the acid?

A. 1.01×10^{-5}
B. 1.01×10^{-7}

C. 9.99×10^{-4}
D. 9.99×10^{-5}

12. A system does 15 J of work and no other changes take place. Which of the following function has fallen by 15 J?

A. Internal energy
B. Specific heat

C. Enthalpy
D. Entropy

13. Which of the following is **CORRECTLY** stated about the spontaneity of a reaction?
- ☒ A. For an endothermic reaction with negative reaction entropy, the reaction is spontaneous at all temperature.
 - B. For an exothermic reaction with positive reaction entropy, the reaction is non-spontaneous at all temperature.
 - C. For an endothermic reaction with positive reaction entropy, the reaction is non-spontaneous at low temperature.
 - D. For an exothermic reaction with negative reaction entropy, the reaction is spontaneous at high temperature.
14. In thermodynamics: solutions in a beaker, a gas in a cylinder, a reaction vessel, a biological cell are examples of a
- ☒ A. thermodynamic variables
 - B. state function
 - C. path function
 - D. thermodynamic system
15. Given the following:
- | | bond energies (kJ/mol) |
|-----|------------------------|
| H-H | 436 |
| N-N | 193 |
| N=N | 409 |
| N≡N | 942 |
- and the enthalpy for the reaction $\frac{1}{2}\text{N}_2(\text{g}) + \frac{3}{2}\text{H}_2(\text{g}) \rightarrow \text{NH}_3(\text{g}); \Delta H = -45 \text{ kJ/mol}$
What is the bond energy for NH_3 ?
- A. 1125
 - B. 1170
 - C. 960
 - D. 1080

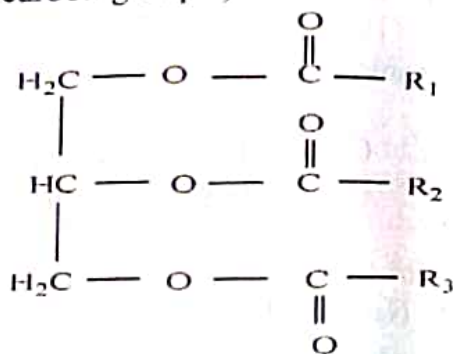
16. The rate law for: $(\text{CH}_3)_3\text{C}-\text{Br} (\text{aq}) + \text{H}_2\text{O} (\text{l}) \rightarrow (\text{CH}_3)_3\text{C}-\text{OH} (\text{aq}) + \text{H}^+ + \text{Br}^- (\text{aq})$ is $\text{rate} = k [(\text{CH}_3)_3\text{C}-\text{Br}]$.

The accepted mechanism for the reaction is

- I. $(\text{CH}_3)_3\text{C}-\text{Br} (\text{aq}) \rightarrow (\text{CH}_3)_3\text{C}^+ (\text{aq}) + \text{Br}^- (\text{aq})$ [slow]
- II. $(\text{CH}_3)_3\text{C}^+ (\text{aq}) + \text{H}_2\text{O} (\text{l}) \rightarrow (\text{CH}_3)_3\text{C}-\text{OH}_2^+ (\text{aq})$ [fast]
- III. $(\text{CH}_3)_3\text{C}-\text{OH}_2^+ (\text{aq}) \rightarrow \text{H}^+ + (\text{CH}_3)_3\text{C}-\text{OH} (\text{aq})$ [fast]

Why **DOES NOT** H_2O appear in the rate law? Because

- A. water is a catalyst in the reaction.
 - B. water is an intermediate species.
 - C. water exists in the liquid state in the reaction.
 - D. water is missing in the rate determining step.
17. Which type of substances has the following general formula? (R_1 , R_2 and R_3 hydrocarbon groups.)



- A. Carboxylic acids
- B. Fats
- C. Alcohols
- D. Fatty acids

18. The rate law for this reaction $\text{CO (g)} + \text{NO}_2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)} + \text{NO (g)}$ is $\text{rate} = k [\text{NO}_2]^2$. Which one of the following could be a plausible mechanism for the reaction?

- A. $\text{NO}_2 \text{ (g)} + \text{NO}_2 \text{ (g)} \rightarrow \text{NO}_3 \text{ (g)} + \text{NO (g)}$ [slow]
 $\text{NO}_3 \text{ (g)} + \text{CO (g)} \rightarrow \text{NO}_2 \text{ (g)} + \text{CO}_2 \text{ (g)}$ [fast]
- B. $\text{NO}_3 \text{ (g)} + \text{NO (g)} \rightarrow \text{NO}_2 \text{ (g)} + \text{NO}_2 \text{ (g)}$ [slow]
 $\text{NO}_3 \text{ (g)} + \text{CO (g)} \rightarrow \text{NO}_2 \text{ (g)} + \text{CO}_2 \text{ (g)}$ [fast]
- C. $\text{NO}_3 \text{ (g)} + \text{NO (g)} \rightarrow \text{NO}_2 \text{ (g)} + \text{NO}_2 \text{ (g)}$ [slow]
 $\text{NO}_2 \text{ (g)} + \text{CO}_2 \text{ (g)} \rightarrow \text{NO}_3 \text{ (g)} + \text{CO (g)}$ [fast]
- D. $\text{NO}_2 \text{ (g)} + \text{NO}_2 \text{ (g)} \rightarrow \text{NO}_3 \text{ (g)} + \text{NO (g)}$ [fast]
 $\text{NO}_3 \text{ (g)} + \text{CO (g)} \rightarrow \text{NO}_2 \text{ (g)} + \text{CO}_2 \text{ (g)}$ [slow]

19. The speed with which a solute goes into solution is

- A. degree of solvation
B. heat of solution
C. rate of dissolution
D. extent of dissolution

20. Which one of the following explains how 2 L of 1 M NaOH aqueous stock solution is converted into a 0.25 M NaOH solution?

- A. Adding 6 L water to the stock solution
B. Adding 8 L water to the stock solution
C. Evaporating 0.25 L of water from the stock solution
D. Evaporating 1 L of water from the stock solution

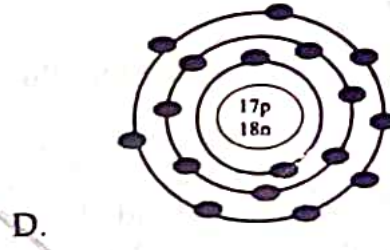
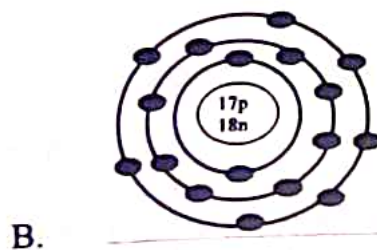
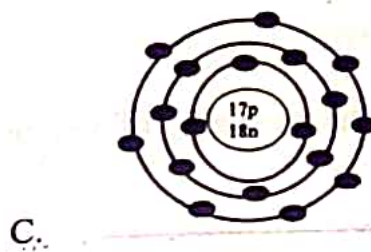
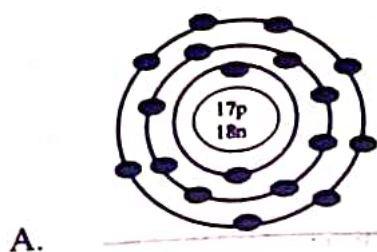
21. Which one of the following **IS TRUE** about isotopes?

- A. Isotopes of an element are atoms that have different number of electrons and protons.
B. Isotopes of an element are identical in mass number.
C. Isotopes of the same elements contain the same number of neutrons.
D. Isotopes of an element have the same chemical properties.

22. Which one of the following describes the quantum mechanical model?

- ☒ A. The locations of electrons in an atom are described in terms of probability.
- B. An atom is represented by a solid indestructible sphere.
- C. An atom is a sphere with a positively charged matter, in which electrons are embedded.
- D. Electrons moved around the nucleus of an atom in circular orbits.

23. Which one of the following is the diagrammatic representation of $^{35}_{17}\text{Cl}$?



24. Which one of the following is an ionic compound?

- ☒ A. CaCl_2
- B. CO_2
- C. PCl_3
- D. H_2O_2

25. An element is located in Group VIIA and 3rd period of the modern periodic table. On the basis of this the element;

- ☒ A. generally forms an ion with -2 charges.
- B. has very low ionization energy.
- C. forms an acidic oxide.
- D. has extremely low electron affinity.

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26. Which of the properties of solvents **WOULD NOT** be affected by the formation of a solution with a non-volatile solute?
- ~~A.~~ Boiling Point C. Chemical identity
B. Chemical potential D. Freezing Point
27. The partial pressure of acetylene, C_2H_2 , in a container is 10 atm at $25^\circ C$. What is the solubility in molarity of C_2H_2 , assume that C_2H_2 obeys Henry's law? ($K_H = 1.04 \text{ mol/L atm}$)
- A. 9.62 C. 104
B. 0.104 ~~D.~~ 10.4
28. What is the normality of a solution that contains 19.6 g of H_2SO_4 in 20 mL of the solution?
- A. 0.02 ~~C.~~ 20
B. 0.01 D. 10
29. Which one of the following illustrates an extensive property?
- A. The electrode potential at standard condition for the half-reaction $2Cu^{2+} + 4e^- \rightarrow 2Cu$ is the same as for $Cu^{2+} + 2e^- \rightarrow 2Cu$.
- B. If a system is divided into parts and it is found that the property of the complete system has a value that is the sum of the values of the property of all the parts.
- C. Doubling the volume of a sample also doubles the mass; so the ratio of mass to volume remains the same.
- ~~D.~~ If water in two beakers at the same temperature are mixed to make a single quantity of water and the temperature of the combined quantity of water is the same as it was in the two separate beakers.
30. The point in a titration at which the number of moles of OH^- ions added to a solution is equal to the number of moles of H^+ ions originally present is
- A. Zero point C. Triple point
B. End point ~~D.~~ Equivalence point

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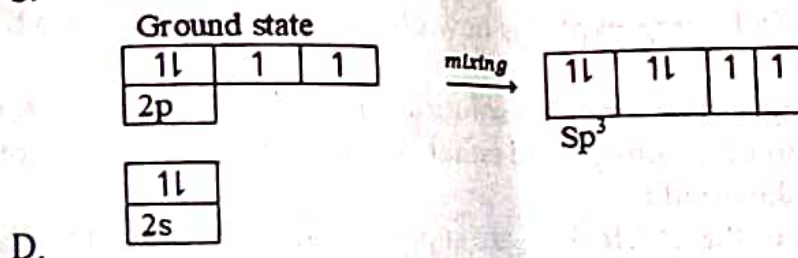
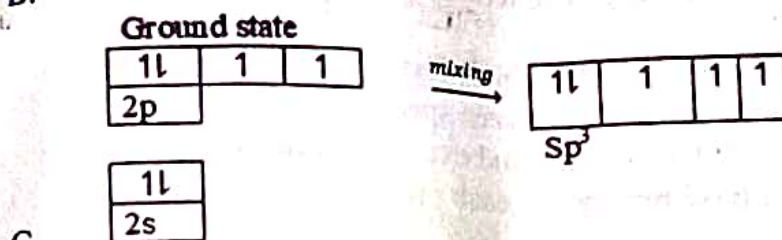
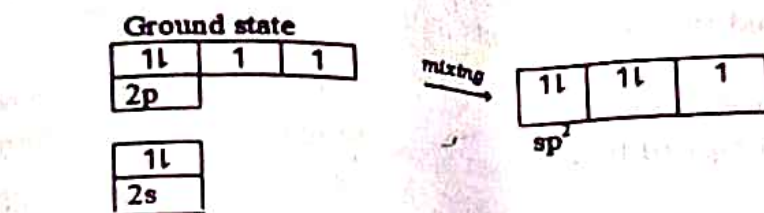
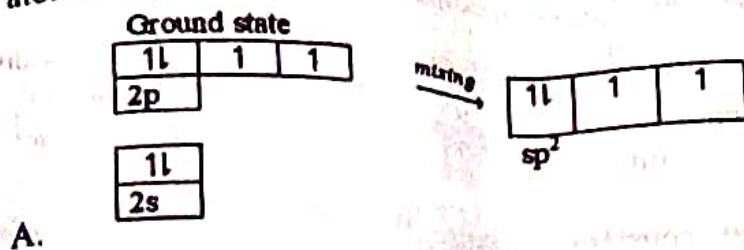
31. Which one of the following is **TRUE** about the electrolysis of a molten hypothetical electrolyte, MX?
- A. X is produced at the negative electrode.
 - B. M is produced at the anode.
 - C. Electrons are not transferred from one species to the other.
 - D. The overall reaction is redox reaction.
32. During the extraction of an unknown metal, the metallic ore was fused with NaOH in air and then dissolved in water to obtain the soluble unknown metal containing species. Reduction with carbon is followed after the solution is treated with HCl. Finally the obtained oxide is reduced to its metallic form by aluminum. The most probable unknown metal would be
- | | |
|-------|-------|
| A. Ca | C. Cr |
| B. Zn | D. Pb |
33. Why tin does **NOT** react with cold water but calcium reacts with cold water quite readily?
- A. Due to its less reactive nature.
 - B. Due to the formation of a thin oxide layer.
 - C. Due to its physical state.
 - D. Due to the pH of water.
34. Why is vulcanization performed in natural rubber?
- A. To convert the monomer unit in natural rubber in to polymer.
 - B. To establish cross-links between polymer chain in the natural rubber.
 - C. To decrease the number of monomer unit in the natural rubber.
 - D. To increase the number of monomer unit in the natural rubber.

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35. Which one of the following explains the determination of the cell potential of Zn-Cu cell using voltmeter?
- A. In the Cu half- cell, Cu metal will be dipped in ZnSO_4 solution.
 - B. In the Zn half- cell, Zn metal will be dipped in CuSO_4 solution.
 - C. Each U-tube served as a salt bridge that can be filled with saturated KCl.
 - D. After the experiment the mass of zinc is increased.
36. Which one of the following is an example of molecular crystals?
- | | |
|-----------------------|-------------------------|
| ~~A.~~ Diamond | C. Solid carbon dioxide |
| B. Sodium Chloride | D. Magnesium |
37. Which one of the following **COULD NOT** be considered as factor influencing the rate of chemical reaction?
- ~~A.~~ Presence of catalyst
 - B. Surface area of reactants
 - ~~C.~~ The rate constant of the reaction
 - D. Temperature of reactants

38. Which one of the following shows the hybridization scheme of the central atom in H_2O ?



39. In which of the following order of a reaction, a plot of $\ln [A]$ versus time is a straight line whose slope is $-k$?

A. First
B. Zero

C. Third
D. Second

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40. A forward and reverse reaction continues to take place at equal rate in microscopic level without change in composition at macroscopic level. This statement is the definition of
- A. Chemical kinetics
B. Dynamic equilibrium
C. Degree of freedom
D. The law of mass action
41. The heat change at constant pressure that results when 1 mole of a compound is formed from its elements in their standard states is
- A. standard entropy of reaction
B. standard enthalpy of reaction
C. standard reduction potential
D. standard enthalpy of formation
42. What is a redox reaction? It is a reaction in which
- A. two insoluble ions in compounds form an insoluble products.
B. electrons are transferred between species.
C. atoms or ions of two compound exchange bonding partners.
D. an acid and a base neutralize each other.
43. Which one of the following explains how electroplating takes place?
- A. The electrolyte is an aqueous solution of a salt of the plated metal.
B. The object to be electroplated either metal or graphite-coated plastic constitutes the anode.
C. The cations in the electrolyte are supplied either by the added salt or from oxidation of the anode.
D. Metal is deposited on the anode by reduction of ions in the electrolyte solution.

44. How does the equivalent mass of a base is obtained?

- A. By multiplying its molar mass by the number of OH^- ions furnished by one formula unit of a base.
- ☒ B. By dividing its molar mass by the number of OH^- ions furnished by one formula unit of a base.
- C. By multiplying its number of equivalent by the number of OH^- ions furnished by one formula unit of a base.
- D. By dividing its number of equivalent by the number of OH^- ions furnished by one formula unit of a base.

45. Which one of the following reactions would occur spontaneously in an aqueous solution at 25°C ? Assume that the initial concentrations of the dissolved species are all 1.0 M.

Half reaction	E° (V)
$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}(\text{s})$	-0.14
$\text{Br}_2(\text{l}) + 2\text{e}^- \rightarrow 2\text{Br}^-(\text{aq})$	+1.07
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag}(\text{s})$	+0.80
$\text{Ni}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Ni}(\text{s})$	-0.25
$\text{Ca}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Ca}(\text{s})$	-2.87
$\text{Cd}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cd}(\text{s})$	-0.40
$\text{Cu}^{2+}(\text{aq}) + \text{e}^- \rightarrow \text{Cu}^+(\text{aq})$	+0.15
$\text{Fe}^{3+}(\text{aq}) + \text{e}^- \rightarrow \text{Fe}^{2+}(\text{aq})$	+0.77

- A. $2\text{Ag}(\text{s}) + \text{Ni}^{2+}(\text{aq}) \rightarrow 2\text{Ag}^+(\text{aq}) + \text{Ni}(\text{s})$
- B. $2\text{Br}^-(\text{aq}) + \text{Sn}^{2+}(\text{aq}) \rightarrow \text{Br}_2(\text{l}) + \text{Sn}(\text{s})$
- C. $\text{Cu}^{2+}(\text{aq}) + \text{Fe}^{2+}(\text{aq}) \rightarrow \text{Cu}^+(\text{aq}) + \text{Fe}^{3+}(\text{aq})$
- ☒ D. $\text{Ca}(\text{s}) + \text{Cd}^{2+}(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{Cd}(\text{s})$

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46. Which of the following statements describes chemical industries? It is
- A. an area of chemistry concerned with the speeds, or rates, at which chemical reactions occur.
 - B. an expression for the chemical composition of a compound in terms of the symbols of the atoms of the elements involved.
 - C. a process in which dissolved solute comes out of solution and forms crystals.
 - ☒ D. a firm that involves usage of raw material to produce desired products by using chemical reactions and refining methods.
47. What will be the product if phosphorus reacts with excess amount of oxygen?
- A. H_3PO_3
 - B. H_3PO_4
 - ☒ C. P_4O_{10}
 - D. P_4O_6
48. One of the basic differences between metallic conduction and electrolytic conduction is:
- ☒ A. The conduction of electricity is due to the movement of ions in an electrolytic conduction while the flow of mobile electrons in metallic conduction.
 - B. The charge carriers in metals are cations and anions while mobile electrons are in electrolytic conduction.
 - C. The charge carriers in metals are localized electrons while strongly held ions are in electrolytic conduction.
 - D. The conduction of electricity is due to the movement of ions in metallic conduction while the flow of mobile electrons in electrolytic conduction.
49. Which one of the following gases is released from combustion of a fuel in an industrial machinery causing global warming?
- A. CO
 - ☒ B. CO_2
 - C. NO
 - D. SO_2

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50. The number of digits to the right of the decimal point denotes:

- ☒ A. Significant figures
☐ B. Decimal places
☐ C. Conversion factor
☐ D. Scientific notation

51. Which of the following would occur during the formation of ions?

- ☒ A. Inner most shell electrons are involved.
☐ B. Valence electrons are shared between atoms.
☐ C. The electronic configuration of an atom forming the ion remains unaffected.
☐ D. Valence electrons are transferred from one atom to another.

52. In order to illustrate the structure and properties of a solid, a student put hundred marble balls in to a box that representing metal cations and the empty space between and around the marble representing the electron sea. Which one of the following bonding is represented by this model?

- ☐ A. Covalent bonding
☐ B. Ionic bonding
☒ C. Metallic bonding
☐ D. Dative bonding

53. Which one of the following explains dipole-dipole force?

- ☐ A. A type of intermolecular force that exists between non-polar molecules.
☒ B. The intermolecular attraction between oppositely charged poles of nearby polar molecules.
☐ C. An attractive force resulting from temporary dipoles induced in molecules.
☐ D. An electrostatic attraction between metal ions and valence electrons.

54. Which one of the following reactions creates energy in our bodies during respiration?

- ☒ A. $C_6H_{12}O_6(s) + 6O_2(g) \rightarrow 6CO_2(g) + 6H_2O(l)$
☐ B. $6CO_2(g) + 6H_2O(l) \rightarrow C_6H_{12}O_6(s) + 6O_2(g)$
☐ C. $C_6H_{12}O_6(s) \rightarrow 2C_2H_5OH(l) + 2CO_2(g)$
☐ D. $N_2H_4CO(g) + H_2O(g) \rightarrow 2NH_3(l) + CO_2(g)$

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55. "A reaction that involves the exchange of positive and negative ions of each reactant". This is a definition of a

- ☒ A. Decomposition reaction
☐ B. Combination reaction
☐ C. Double displacement reaction
☐ D. Single displacement reaction

56. What are the four quantum numbers for an electron in a 2s orbital?

- A. 2, 1, 0, $+\frac{1}{2}$
B. 2, 1, 1, $+\frac{1}{2}$
☒ C. 2, 0, 0, $+\frac{1}{2}$
D. 2, 0, 1, $+\frac{1}{2}$

57. A student uses a digital balance (± 0.02), a thermometer (± 0.1), a measuring cylinder (± 0.05), and a beaker (± 0.5) during an experiment. Which of the above instrument is more precise?

- A. The thermometer
B. The digital balance
☒ C. The Beaker
☐ D. The measuring cylinder

58. Which one of the following explains that light has **BOTH** wave and particle nature?

- A. A beam of light is a stream of particles of electromagnetic energy, with energy proportional to the observed frequency of the light.
☒ B. An electron that is bound to the nucleus behaves like a standing wave and its position is given in terms of probability.
C. Microscopic systems, such as atoms and molecules can possess wave properties.
D. Light consists of electromagnetic waves which have an oscillating electric and magnetic disturbance that spreads electromagnetic radiation.

59. A nuclear breakdown in which particles or electromagnetic radiation is emitted is _____.

- A. Radioactive isotopes
B. Radioactivity
☒ C. Radio wave
☐ D. Radioactive decay

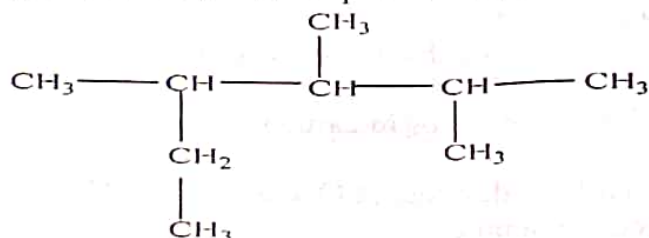
60. Which statement describes valence shell electron pair repulsion (VSEPR) theory?
- A. Each degenerate orbitals are occupied by a single electron before the second electrons of opposite spin enters the orbitals.
 - B. Electron pairs will be as far apart from each other in three dimensional spaces as possible.
 - C. Atomic orbitals are capable of combining to produce new electron distribution.
 - D. A covalent bond is formed when orbitals of two atoms overlap and the overlap region is occupied by a pair of electrons.

61. Which one of the following is a colorless, flammable volatile liquid and miscible with non-polar solvents?

A. Benzene
B. Ethanol

C. Methanol
D. Methane

62. The IUPAC name of the compound with the following structure is:



- A. 3, 4, 5-Trimethylhexane
 - B. 2-Ethyl-3, 4-dimethylpentane
 - C. 4-Ethyl-2, 3-dimethylpentane
 - D. 2, 3, 4-Trimethylhexane
63. Which one of the following petroleum products is used as a solvent and in dry cleaning?
- A. kerosene
 - B. Petroleum ether
 - C. Paraffin wax
 - D. Diesel oil

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64. Which one of the following is **NOT TRUE** about laboratory preparation of ethylene by dehydration of ethanol?

- ☒ A. The collected ethylene gas in the experiment decolorizes Br_2 in CCl_4 .
- B. Concentrated sulphuric acid is used as a dehydrating agent.
- C. Combustion of ethylene produces a non-luminous flame.
- D. The collected ethylene gas in the experiment decolorizes alkaline KMnO_4 .

65. Which one of the following is an essential nutrient of plants?

- | | |
|-------------|---|
| A. Aluminum | C. Sodium |
| B. Lead | ☒ D. Potassium |

66. Which one of the following is the main reason for the large difference in dipole-moment between the triatomic molecules water (H_2O) and carbon dioxide (CO_2)?

- A. The presence of lone pairs of electrons in water (H_2O) partially reinforces bond polarities.
- ☒ B. The higher electronegativity difference among atoms in water (H_2O) than in carbon dioxide (CO_2).
- C. The presence of multiple bonds in carbon dioxide (CO_2) while single bonds in water.
- D. The bonds in carbon dioxide (CO_2) are directed 180° and bond polarity reinforces each other.

67. How many non-bonding pairs of electrons are present in the valence shell of the central atom in ClF_3 ?

- | | |
|---------------------------------------|------|
| ☒ A. 1 | C. 3 |
| B. 0 | D. 2 |

68. Which one would be observed in a laboratory activity that is applied on naphthalene?

- A. It is soluble in benzene.
- B. It conducts electricity in molten state.
- ☒ C. It shows relatively high melting point.
- D. It is soluble in water.

69. Which one of the following statements is CORRECT?

- ☒ A. Valence bond model simply shows the linkage between atoms and the presence of lone pairs.
- B. Lewis model predicts the shapes of molecules based on electron-pair repulsions around a central atom.
- C. In Lewis model, electrons repel one another and a pair of bonding electrons lies far apart as possible.
- D. Valence bond model explain molecular shape from the interaction of atomic orbitals.

70. The energy required to completely separate one mole of a solid ionic compound into gaseous ions is a/an

- A. Ionization energy
- B. Activation energy
- ☒ C. Lattice energy
- D. Bond energy

71. Which of the following compounds reacts with basic oxide to form salt and water?

- A. Acid
- B. Acidic oxide
- C. Base
- D. Neutral oxide

72. Which one of the following reaction represents formation of an acid from its salt?

- A. $\text{CH}_3\text{COOH (aq)} + \text{CH}_3\text{OH (aq)} \rightarrow \text{CH}_3\text{CO}_2\text{CH}_3 \text{ (aq)} + \text{H}_2\text{O (l)}$
- B. $\text{P}_2\text{O}_5 \text{ (s)} + 6\text{NaOH (aq)} \rightarrow 2\text{Na}_3\text{PO}_4 \text{ (s)} + 3\text{H}_2\text{O (l)}$
- ☒ C. $\text{NaNO}_3 \text{ (s)} + \text{H}_2\text{SO}_4 \text{ (l)} \rightarrow \text{NaHSO}_4 \text{ (s)} + \text{HNO}_3 \text{ (l)}$
- D. $\text{AgNO}_3 \text{ (aq)} + \text{NaCl (aq)} \rightarrow \text{NaNO}_3 \text{ (aq)} + \text{AgCl (s)}$

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73. The negative logarithm of the molar concentration of hydrogen ion to the base ten is known as _____.

- ☒ A. pH
- ☐ B. K_w
- ☐ C. K_a
- ☐ D. pOH

74. Which one of the following galvanic cell is an example of secondary cell?

- ☐ A. Daniel's cell
- ☐ B. Lead storage cell
- ☒ C. Leclanche dry cell
- ☐ D. Zinc-carbon dry cell

75. Which one of the following is responsible for electrolytic conductivity?

- ☐ A. Ions strongly held in a fixed position.
- ☐ B. Free electrons.
- ☒ C. Electrons held in a fixed position.
- ☐ D. Movement of ions.

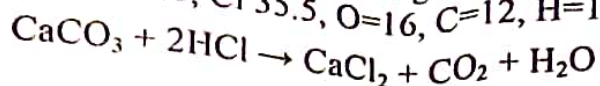
76. Which one of the following statement is **TRUE** about gases?

- ☐ A. The particles of gases are closely packed together.
- ☒ B. The volume of a gas is kept constant as pressure increases.
- ☐ C. When a gas sample is heated at constant pressure, its volume decreases.
- ☐ D. The volume of a gas is equal to the volume of its container.

77. Which one of the following mathematical expressions is derived from the combination of Boyle's law, Charles' law and Avogadro's law?

- ☒ A. $V \propto \frac{nT}{P}$
- ☐ B. $V \propto \frac{T}{nP}$
- ☐ C. $V \propto \frac{nP}{T}$
- ☐ D. $V \propto \frac{P}{nT}$

78. 180 g of calcium carbonate (CaCO_3) is allowed to react with 156 gram of hydrochloric acid (HCl), according to the following reaction. Which of the following substances is the limiting reagent?
(Atomic masses; $\text{Ca}=40$, $\text{Cl}=35.5$, $\text{O}=16$, $\text{C}=12$, $\text{H}=1$)



- A. HCl
~~B. CaCO_3~~
C. H_2O
D. CO_2
79. Which one of the following isomers differs in the arrangement of carbon atoms in the longest continuous carbon chain?
- A. Geometric
B. Position
C. Cis-trans
~~D. Structural~~
80. Which one of the following methods can be used in the laboratory for the preparation of alkynes?
- A. Dehydration of alcohols with concentrated sulphuric acid.
B. Heating of sodium salt of an organic acid with soda lime.
~~C. Dehydrohalogenation of adjacent dihalides with sodium hydroxide.~~
D. Hydrogenation of alkenes with a metal catalyst.

The End 

