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QUESTION 2

Fill in the blanks with the best term from the Word Bank.

Word Bank			
high	low	strongly	weakly

Ionic compounds have high melting points. This property is due to the fact that positive and negative ions are strongly attracted to each other.

QUESTION 3

Select the best answer.

Which of the following **best** describes the arrangement of ions in an ionic compound?

- The ions are widely spaced because similar charges repel each other.
- The ions are closely packed because opposite charges repel each other.
- **The ions are closely packed because opposite charges attract each other.**
- The ions are widely spaced because similar charges attract each other.

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QUESTION 5

Select the best answer from the drop-down menu to complete the sentence.

Ionic compounds have rigid ✓ crystal lattices because of the strong attraction between ions.



QUESTION 7

Select the best answer.

The crystals of an ionic compound are a bright orange color. What does this **most likely** tell you about the ionic compound?



- ○ The compound contains a halogen.
- ○ The compound contains an alkali metal.
- ◉ The compound contains a transition element. ✓
- ○ The compound contains a metalloid.

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QUESTION 9

Drag and drop the best phrase to complete the sentence.

The crystal lattice describes a [three-dimensional and regular ✓] arrangement of ions in a crystal structure.

two-dimensional and regular

two-dimensional and irregular

three-dimensional and irregular

QUESTION 11

Drag and drop the best terms to complete the sentences.

Some ionic compounds can [conduct ✓] electricity in an aqueous solution.

For this reason, ionic compounds are called [electrolytes ✓].

do not conduct

sometimes conduct

nonelectrolytes

neutral

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QUESTION 13

Select the best answer.

One of the factors scientists use to classify ionic compounds in minerals is their _____.



- (●) anions ✓
- ( ) cations
- ( ) nonmetals
- ( ) metals

QUESTION 15

Select the best answer from the drop-down menu to complete the sentence.

A large amount of ✓ energy for ionic compounds to change to a different state of matter.

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QUESTION 17

Select the best answers.

Select the ionic compounds based on their properties.

- [x] A substance that has a melting point of 808°C and conducts electricity in an aqueous solution. ✓
- [ ] A substance that has white crystals that are easily broken and do not conduct electricity.
- [ ] A substance with blue crystals that do not conduct electricity when dissolved in water.
- [x] A substance that has bright orange crystals that are hard, rigid, and brittle. ✓
- [ ] A substance with yellow crystals that are soft and a melting point of −6°C.
- [ ] A substance that has orange crystals that are flexible and melt at 15°C.

QUESTION 19

Select the best answer.

Sodium chloride splits completely into its positive and negative ions when placed in water. Which statement accurately describes sodium chloride?

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- Sodium chloride is a strong electrolyte and conducts electricity poorly.
- ◉ Sodium chloride is a strong electrolyte and conducts electricity efficiently. ✓
- Sodium chloride is a weak electrolyte and conducts electricity efficiently.
- Sodium chloride is a weak electrolyte and conducts electricity poorly.

QUESTION 21

Select the best words from the drop-down menus to complete the sentence.

One of the ways minerals are named is through their anions. Anions in carbonate compounds are a combination of [carbon ✓] and oxygen, while ionic compounds containing boron ions are called [borates ✓].

QUESTION 23

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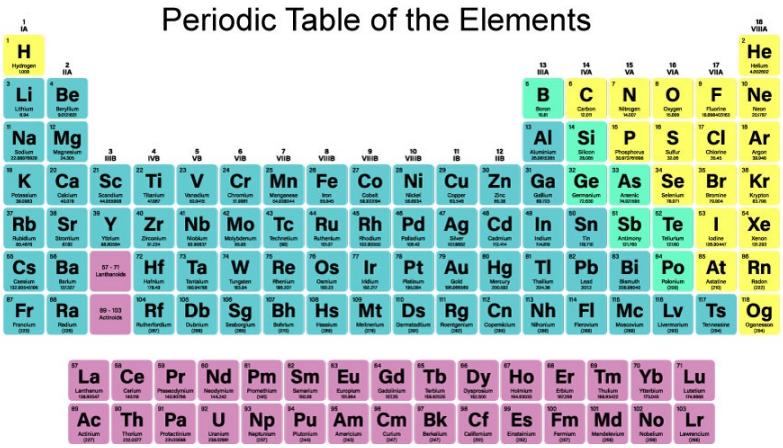
Select the best answers.

Which of the following ionic compounds are **most likely** to be brightly colored?





- [ ] sodium bromide
- [ ] sodium chloride
- [ ] potassium nitrate
- [x] copper(II) chloride
- [x] nickel(II) sulfate
- [ ] lithium sulphate

QUESTION 25

Select the best answers.

Which of the following statements are correct about electrical conductivity in ionic compounds?

- [x] Ionic compounds conduct electricity in a liquid form.
- [ ] Ionic compounds conduct electricity only if they are halides.
- [ ] Ionic compounds never conduct electricity.
- [x] Ionic compounds dissolved in solutions conduct electricity.
- [ ] Ionic compounds only conduct electricity as a solid.
- [ ] Ionic compounds conduct electricity in a crystal lattice structure.

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QUESTION 27

Select the best answers.

How can an ionic compound be made to conduct electricity?

- [ ] pass electric current into the solid ionic compound
- [x] dissolve the ionic compound in water
- [ ] vaporize the ionic compound to become a gas

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- [ ] mix together crystals from two different ionic compounds
- [ ] crush the ionic compound into dust
- [x] melt the ionic compound ✓

QUESTION 29

Select the best answers.

Which of the following statements describe ionic compounds?

- [x] substances that do not conduct electricity in the solid state ✓
- [ ] substances that do not conduct electricity when dissolved in water
- [ ] substances that can be hammered into thin sheets
- [ ] substances which are usually gases at room temperature
- [x] substances that conduct electricity when dissolved in water ✓
- [ ] substances that cannot conduct electricity in molten form

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