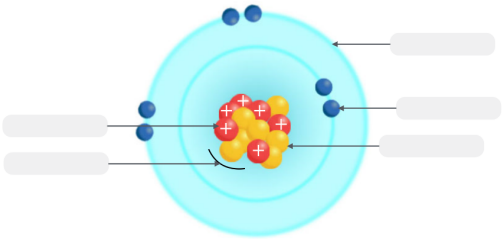
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QUESTION 1

0:00 / 0:04

Drag and drop the terms to correctly label the atomic model.



proton ✓

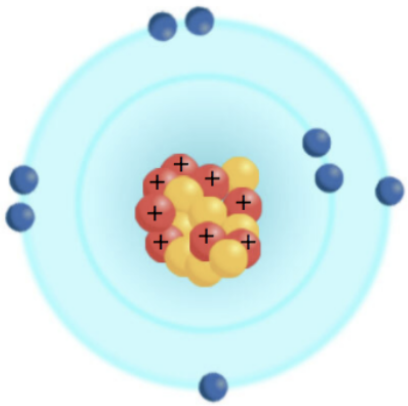
nucleus ✓

QUESTION 2

Select the correct answer.

0:00 / 0:08

Examine the atomic model. Identify the number of protons, neutrons, and electrons in the given atom.



- 6 protons, 1 neutron, 6 electrons
- 9 protons, 10 neutrons, 9 electrons
- 1 proton, 0 neutrons, 1 electron
- 8 protons, 8 neutrons, 8 electrons ✓

QUESTION 3

Drag and drop the correct choice to complete the sentence.

0:00 / 0:04

Compared to the charge of an electron, the proton charge is equal and of opposite sign ✓.

equal and of the same sign | smaller and of opposite sign | greater and of the same sign

QUESTION 4

Drag and drop the best choices to complete the sentences.

0:00 / 0:12

Protons ✓ are the positively charged particles of an atom.

Neutrons ✓ are atomic particles that have no charge.

Electrons ✓ are the negatively charged particles of an atom.

Energy levels | Nuclei

QUESTION 5

Select the correct choice.

0:00 / 0:21

The masses of the three subatomic particles can be expressed in amu (atomic mass units). The table below shows the mass of the different subatomic particles. Examine the table to answer the question that follows.

Subatomic Particle	proton	neutron	electron
Mass (amu)	1	1	0.0005

Where is most of the mass of an atom located?

Most of the mass of an atom is in the energy level closest to the nucleus.

Most of the mass of an atom is in the outermost energy level.

Most of the mass of an atom is in the nucleus. ✓

Most of the mass of an atom is in the energy levels.

QUESTION 6

Select the best choice.

0:00 / 0:17

The nucleus of an atom contains positively charged protons and neutrons. The atoms also contain negatively charged electrons that are found in energy levels outside the nucleus. If atoms contain charged particles, why do they not have a charge?

Atoms contain the same number of protons as neutrons.

Atoms contain charged particles only in the nucleus.

Atoms contain an unequal number of protons and electrons.

Atoms contain the same number of protons as electrons. ✓

QUESTION 7

Select the correct choice.

0:00 / 0:02

Which of the following is the smallest?

a cell

an atom ✓

sand particle

a bacterium

QUESTION 8

Select the correct choice.

0:00 / 0:11

An ion is a charged atom. An ion can have a positive or a negative charge. What is structurally different between an ion and an atom?

An ion has electrons in its nucleus. An atom has electrons moving in energy levels.

An atom has a different number of electrons than protons. An ion always has the same number of protons as electrons.

An ion has a different number of electrons than protons. An atom always has the same number of protons as electrons. ✓

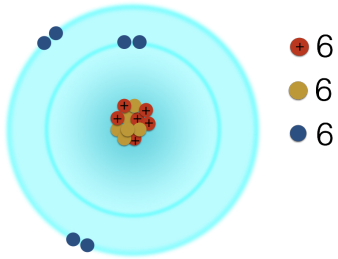
An ion does not have a nucleus. An atom has a nucleus.

QUESTION 9

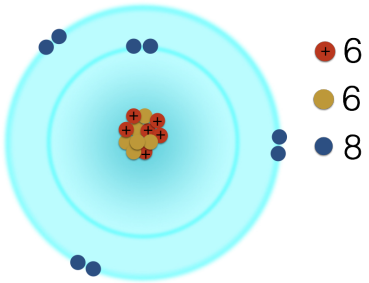
Select the correct choice.

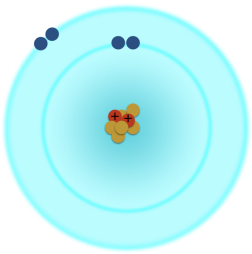
0:00 / 0:14

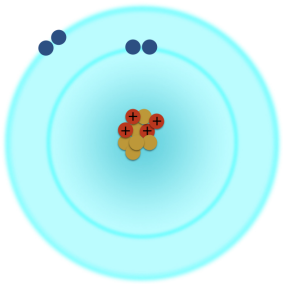
The mass number of an atom is the total number of neutrons and protons. A neutral atom has 6 electrons and 6 neutrons. Select the **CORRECT** atomic model that represents the described atom.



✓





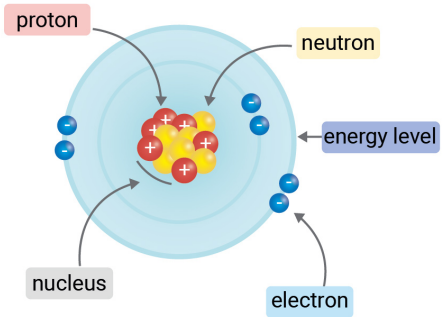


QUESTION 10

Select the correct choice.

0:00 / 0:09

The figure represents a model of a carbon atom. Which of the following is **NOT** correct about the carbon atom model shown in the figure?



The number of protons is equal to the number of electrons.

The carbon atom has six protons.

The model represents a charged atom. ✓

The model represents a neutral atom.

QUESTION 11

Select the correct choice.

0:00 / 0:03

Which of the following choices do **NOT** match?

proton: positive charge: energy level ✓

neutron: no charge: nucleus

proton: positive charge: nucleus

electron: negative charge: energy level

QUESTION 12

Select the correct choice.

0:00 / 0:15

An ion is a charged atom. An ion has an unequal number of protons and electrons. A fluoride ion has a charge of -1.

Which of the following is **CORRECT** about the fluoride ion?

The fluoride ion has one more proton than electrons.

The fluoride ion has two more electrons than protons.

The fluoride ion has an extra neutron.

The fluoride ion has one more electron than protons. ✓

QUESTION 13

Select the correct choice.

0:00 / 0:06

Select the statement that **best** describes the atom if it had more electrons than protons.

- It would have a positive charge.
- ✓ It would have a negative charge.
- It would have the same number of neutrons and electrons.
- It would have no electric charge.

QUESTION 14

Select the correct choice.

0:00 / 0:08

Matter is anything that has mass and takes up space. The _____ is the smallest unit of all matter.

- compound
- ✓ atom
- cell
- organism

QUESTION 15

Select the correct choice.

0:00 / 0:03

Which of the following statements **best** describes an atom?

- An atom is composed of an electron level surrounded by a nucleus.
- An atom is composed of a nucleus containing electrons, and energy levels containing protons.
- An atom is composed of a negatively charged nucleus, surrounded by energy levels.
- An atom is composed of a nucleus containing protons and neutrons, and energy levels containing electrons. ✓